

AP CHEMISTRY CHAPTER 6 PRACTICE TEST

- For which chemical equation will the heat of reaction  $\Delta H^\circ$ , correspond to the molar heat of formation of liquid water?
  - $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
  - $\text{H}_{(g)} + \text{H}_{(g)} + \text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
  - $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)}$
  - $\text{H}^+_{(aq)} + \text{OH}^-_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)}$
- For the process  $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$ ,  $\Delta H = -247$  kJ. The numeric value of -242 kJ represents heat energy that is
  - absorbed by the reaction
  - required to break the H-H and O=O bond
  - stored as potential energy
  - released by the reaction
- An 11.0 g sample of propane is placed in a bomb calorimeter with excess oxygen and is ignited. The initial temperature of the calorimeter is 25.00°C and its total heat capacity is 97.1 kJ/°C. The reaction raises the temperature of the calorimeter is 30.75°C. The approximate molar heat of combustion for propane is
  - 50.6 kJ/mol
  - 558 kJ/mol
  - $-2.24 \times 10^3$  kJ/mol
  - $-1.20 \times 10^4$  kJ/mol
- What is the change in enthalpy at 25°C and 1 atm for the production of 3.00 mol  $\text{SnO}_{(s)}$ ?
 
$$\text{Sn}_{(s)} + \text{SnO}_{2(s)} \rightarrow 2\text{SnO}_{(s)} \quad \Delta H^\circ = 16.2 \text{ kJ}$$
  - 24.3 kJ
  - 16.2 kJ
  - 5.40 kJ
  - 16.2 kJ
  - 24.3 kJ
- How much heat is liberated at constant pressure if 0.896 g of calcium carbonate reacts with 92.3 mL of 0.584 M hydrochloric acid?
 
$$\text{CaCO}_{3(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{2(aq)} + \text{H}_2\text{O}_{(l)} + \text{CO}_{2(g)} \quad \Delta H^\circ = -15.2 \text{ kJ}$$
  - 0.410 kJ
  - 8.88 kJ
  - 13.6 kJ
  - 0.546 kJ
  - 0.136 kJ

# Ap Chemistry Chapter 6 Test

**Richard H. Langley, John Moore**



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