

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

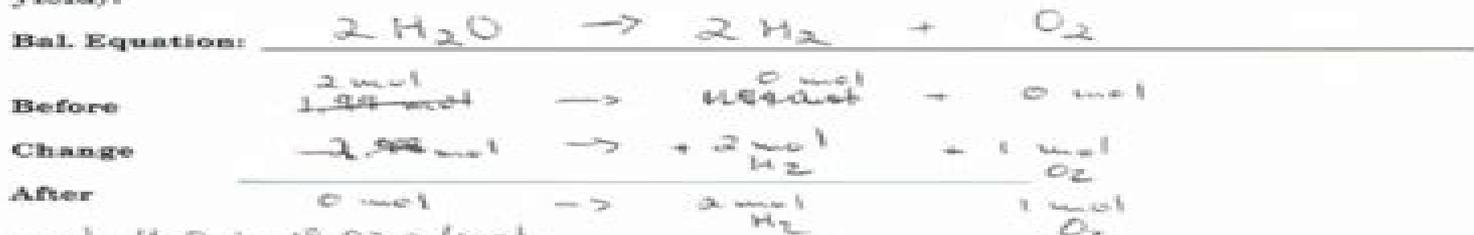
↳ from B-C-A chart

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



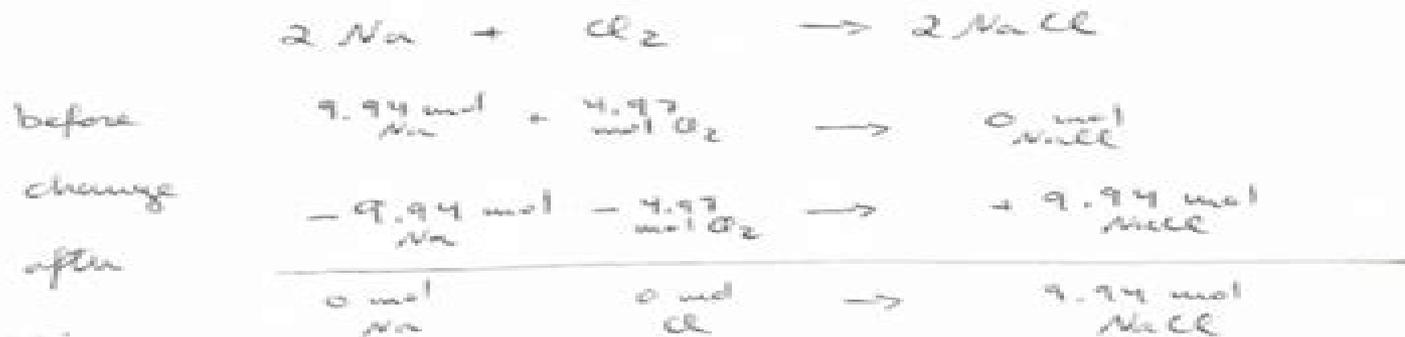
1 mol H₂O : 18.02 g/mol

36 g $\times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g}} = 1.99 \text{ mol H}_2\text{O}$

1 mol H₂ : 2.02 g

2 mol H₂ = 4.04 g H₂

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol NaCl = 58.44 g
 $581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g}}$
 $= 9.94 \text{ mol NaCl}$

1 mol Na = 22.99 g

9.94 mol Na \times 22.99 g = 228.52 g Na

Modeling Chemistry U8 V2 Answers

Roman Wölfel



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