

# 1 • Matter and Measurement

*2* *100 BP* *1F* *212* *of degrees are smaller than °C*  
*0 FP* *32*  
**PRACTICE TEST**

- Which physical state of matter exhibits the greatest change in volume with changes in temperature or pressure? *mostly empty space*  
 a) solid      b) liquid      **c) gas**  
*easy to compress*
- According to the Kinetic Molecular Theory particles of a solid  
 a) are bound in a regular array and do not move.  
 b) float freely within an array occupying various positions relative to neighbors. *True*  
 c) have no relationship to the microscopic structure of the solid.  
**d) vibrate back and forth but do not move past immediate neighbors.**  
 e) float freely in the inside but do not move on the surface.
- What volume of a liquid having a density of 1.48 g/cm<sup>3</sup> is needed to supply 5.00 grams of the liquid?  
 a) 0.296 cm<sup>3</sup>      **d) 3.38 cm<sup>3</sup>**  
 b) 1.48 cm<sup>3</sup>      e) 7.40 cm<sup>3</sup>  
 c) 2.26 cm<sup>3</sup>      *5.00g \* 1 cm<sup>3</sup> / 1.48g =*
- The density of aluminum is 2.70 g/cm<sup>3</sup>. If a cube of aluminum weighs 13.5 grams, what is the length of the edge of the cube?  
 a) 5.00 cm      d) 0.312 cm  
**b) 1.71 cm**      e) 0.200 cm  
 c) 1.25 cm  
*13.5g \* 1 cm<sup>3</sup> / 2.70g = 5.00 cm<sup>3</sup>*  
*∛5.00 cm =*

- Which temperature change is the smallest?  
 a) 10°C to 20°C      **d) 10°F to 20°C**  
~~b) 10 K to 20°C~~      **e) 10°F to 20°F**  
 c) 10 K to 20 K
- The number of significant figures in 0.06060 x 10<sup>-5</sup> is  
 a) 2      d) 5  
 b) 3      e) 6  
**c) 4**
- The number, three hundred and fifty thousand, written in scientific notation is best written as  
 a) 350      d) 3.50 x 10<sup>5</sup>  
 b) 3.5 x 10<sup>6</sup>      e) 3.50 x 10<sup>-5</sup>  
**c) 3.5 x 10<sup>5</sup>** *assume least # of digits significant digits*  
*350,000 = 3.5 x 10<sup>5</sup>*
- The mass of a sample weighted on an electronic balance that is sensitive to ±0.3 mg is 1.2300 g. The number of significant figures in this measurement is *.3mg = .0003g*  
 a) 1      d) 4  
 b) 2      **c) 5**      *1.2300 ± .0003*  
 e) 3
- What is the numerical value of:  
 1.5 cm - 7.222 x 10<sup>-1</sup> cm?  
 a) 0.7778 cm      **d) 0.8 cm**  
 b) 0.778 cm      e) 7.072 x 10<sup>-1</sup> cm  
 c) 0.78 cm  

$$\begin{array}{r} 1.5 \text{ cm} \\ - 0.7222 \text{ cm} \\ \hline 0.7778 \text{ cm} \\ \hline 0.8 \text{ cm} \end{array}$$

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**AN Whitehead**



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