

Limitations : (i) During anodic protection method, the corrosion does take place, but at a very slow rate. In other words, corrosion rate is not zero. (ii) This method can be applied only to those metals/alloys which passivate.

Applications : Stainless steel containers used for transporting corrosive chemicals (e.g., concentrated acids) are protected from corrosion by passivating them by maintaining them at proper E_f and I_p .

Short Answer Questions

1. Define corrosion.

(Bombay, May 93, Nov. 97)

Ans. Any process of deterioration and consequent loss of solid metallic materials through an unwanted chemical or electrochemical attack by its environment, starting at its surface, is called corrosion.

2. What is meant by rusting of iron ?

Ans. The atmospheric corrosion of iron and steel, leading to the formation of a layer of reddish scale and powder of oxide (Fe_2O_3) on the surface.

3. What is rust ?

Ans. Hydrated ferric oxide (Fe_2O_3) with $Fe(OH)_3$ - Brown rust ; mixed ferrous and ferric oxides ($Fe_3O_4 = FeO \cdot Fe_2O_3$) - Black rust.

4. What is dry corrosion ?

Ans. Corrosion caused by direct chemical action of atmospheric gases such as oxygen, halogen, sulphur dioxide, hydrogen sulphide with metal surfaces in immediate proximity.

5. Formation of which types of metal oxide film cause rapid and continuous corrosion.

Ans. Volatile oxide film and porous oxide film.

6. Formation of which types of metal oxide film prevents corrosion.

Ans. Fine-grained tightly adhering, impervious oxide film ; and highly unstable oxide film.

7. What is Pilling-Bedworth rule ?

Ans. An oxide film is protective (or non-porous) if the volume of the oxide is at-least as great as the volume of the metal from which it is formed. On the other hand, if the volume of the oxide film is less than the volume of the metal, then the oxide film is porous and non-protective.

8. What is wet corrosion ?

Ans. A corrosion which occurs when a conducting liquid is in contact with : (i) metal or (ii) when two dissimilar metals or alloys are either immersed or dipped partially in a solution.

9. What is electrochemical corrosion ?

Ans. This type of corrosion occurs due to the existence of separate 'anodic' and 'cathodic' areas between which current flows through the conducting solution. At anodic area oxidation occurs and anodic part of metal is destroyed.

10. What is galvanic corrosion ?

Ans. When two dissimilar metals are electrically connected and exposed to an electrolyte, the metal higher in electrochemical series undergoes corrosion.

11. The rate of metallic corrosion increases with increase in temperature. Give reason.

(Bangalore, Aug. 95)

Ans. With increases of temperature of the environment, the rate of reaction as well as rate of diffusion increase, thereby corrosion rate increases.

12. Iron corrodes faster than aluminium, even though iron is placed below aluminium in the electrical series. Why ?

(Bangalore, Jan. 93, Aug. 95)

Ans. This is because aluminium forms a non-porous, very thin, tightly adhering protective oxide film (Al_2O_3) on its surface and this film does not permit corrosion to occur.

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Yijin Wang



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