

7.3

BONDING IN METALS

Section Review

Classroom

- 1. Describe metallic bonding.
- 2. Explain the relationship between the level of delocalized electrons and the strength of metallic bonding.

Worksheet

- 1. Review notes.
- 2. Do HW.

Part A: Completion

Use the following information to complete the sentences. Write your answers in the spaces provided.

1. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
2. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
3. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
4. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
5. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.

- 6. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 7. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 8. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 9. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 10. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 11. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 12. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 13. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 14. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.
- 15. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions.

Part B: True/False

Write whether each statement is true (T) or false (F).

1. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions. **T**
2. The strength of metallic bonding is **strong** because the delocalized electrons are **strongly attracted** to the positive metal ions. **T**

Bonding In Metals Section Review Answers Key

Arshad Iqbal



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