

## Chapter 5 Chapter Review Question Answers

- Describe the scientists' contributions:
  - Cannizzaro—developed a standard method for measuring atomic masses, which allowed chemists to search for periodic trends among elements.
  - Mendeleev—organized elements according to increasing atomic mass and noticed that similar properties appeared periodically.
  - Moseley—discovered that atomic number, not atomic mass, should be the basis of organizing elements on the periodic table.
- The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- Groups of elements exhibit similar chemical and physical properties and behaviors.
- Electron configurations:
  - Generally, the configurations of the outermost electrons shells of elements within the same group are the same. There are a number of exceptions to this rule, however, among the transition elements.
  - Their outer shells are completely filled.
- The length of a period is determined by the total number of electrons that can fill the outer sublevels of the elements of that period.
- An element's period corresponds to its highest occupied main energy level. (s & p)
- Blocks & Groups
  - The type of sublevel being filled in successive elements of that block.
  - s-block: Groups 1 & 2; p-block: Groups 13–18; d-block: Groups 3–12; f-block: No group #s, as they are between groups 3 & 4, but period 6 elements are in the Lanthanide series, and period 7 elements are in the Actinide series.
- Alkali Metals
  - The Group 1 elements, except Hydrogen.
  - Alkali metals are extremely reactive, they react violently with water, they have low densities, they are soft enough to be cut with a knife, and freshly cut alkali metals are very shiny silver colored metals.
- Alkaline Metals
  - The Group 2 elements.
  - They are slightly less reactive than Group 1 elements; they are harder, duller, denser than Group 1 elements. They also have slightly higher melting points than the Group 1 elements.
- Group Configurations:
  - Transition Metals (d-block)
    - Group 3:  $ns^2 (n-1)d^1$
    - Group 4:  $ns^2 (n-1)d^2$
    - Group 5:  $ns^2 (n-1)d^3$
    - Group 6:  $ns^2 (n-1)d^4$  {actually is  $ns^1 (n-1)d^5$  for Cr & Mo}
    - Group 7:  $ns^2 (n-1)d^5$
    - Group 8:  $ns^2 (n-1)d^6$
    - Group 9:  $ns^2 (n-1)d^7$
    - Group 10:  $ns^2 (n-1)d^8$
    - Group 11:  $ns^2 (n-1)d^9$  {actually is  $ns^1 (n-1)d^{10}$  for all group 11 elements}
    - Group 12:  $ns^2 (n-1)d^{10}$
  - The group number is the sum of the s & d electrons.
- The transition elements.
- p-block
  - The p-block consists of nonmetals at the right, metalloids in the middle along the staircase, and metals at the left.
  - The p-block metals are generally harder and denser than the s-block metals, but softer and less dense than the d-block metals.
- Halogens
  - The Group 17 elements.

# Modern Chemistry Chapter 5 Review Answer Key

**R Barnett**



**Modern Chemistry Chapter 5 Review Answer Key:**

## **Modern Chemistry Chapter 5 Review Answer Key** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Modern Chemistry Chapter 5 Review Answer Key**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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