

# Colorimetric Titration Experiment - Determine the amount of vitamin C in a medium peach

A suitable method for the determination of vitamin C (ascorbic acid,  $C_6H_8O_6$ ) is a titration with potassium iodate ( $KIO_3$ ). In this reaction, potassium iodate is used as a titrant and when added to an ascorbic acid solution that contains strong acid and potassium iodide (KI), the potassium iodate reacts with potassium iodide, liberating molecular iodine ( $I_2$ ) as shown in the reaction below:



During the titration, as long as the solution contains ascorbic acid, the  $I_2$  produced in equation 1 is used up in a rapid reaction with ascorbic acid (equation 2), during which dehydroascorbic acid ( $C_6H_6O_6$ ) and iodide ion ( $I^-$ ) are formed:



Once all the ascorbic acid has been consumed, any excess iodine ( $I_2$ ) will remain in solution. This excess iodine reacts with starch, to form an intensely blue colored complex, indicating that the endpoint is reached.

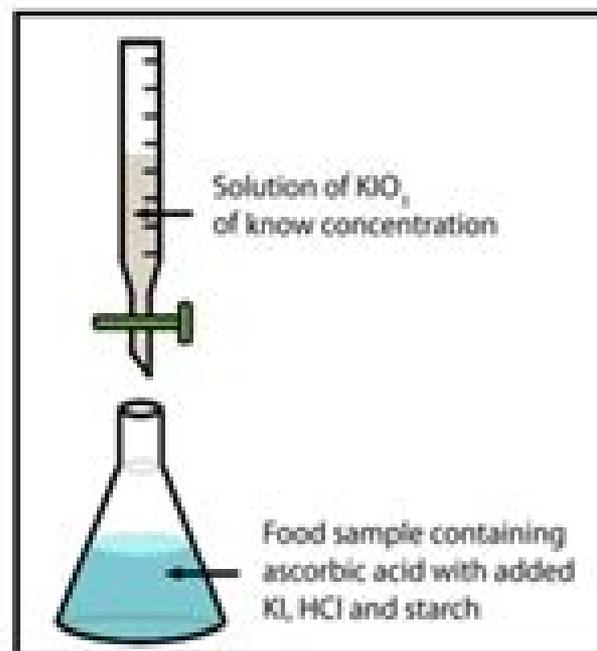
The experimental setup is shown schematically to the right.

The virtual lab stockroom contains the following solutions:

**Solid potassium iodate** - ( $KIO_3$ ). You will need to create a stock solution with this in order to perform your titration. (Note: The molar mass of potassium iodate is 214.001 g/mol)

**Potassium iodide solution** - This is provided as a concentrated solution of KI in water. Since you are required to use an excess KI, the amount does not need to be measured precisely. Use 5 mL of KI solution for each titration.

**2M HCl** - The addition of acid is necessary to provide the acidic conditions required in the iodometric titration (see Equation 1). Use 4mL of the HCl solution in each titration.



# Determination Of Vitamin C Concentration By Titration

**Rachel Sandford**



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