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**Shasha Hu**



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When the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients of the reactants and products is 4 a a q is

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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