

NAME _____ DATE _____ CLASS _____

CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , NO_2 , and NO are 32.00 g/mol, 46.01 g/mol, and 30.01 g/mol, respectively.

Amount of O_2	Amount of NO_2	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO_2	2 molecules O_2
2 molecules	4 molecules	1, 4 molecules	1, O_2	3, 4 molecules NO
1.00 mol	2.00 mol	3, 2.00 mol	3, none	3, none
4.00 mol	6.00 mol	2, 4.00 mol	2, NO_2	3, 2.00 mol O_2
0.00 mol	0.00 mol	10, 7.00 mol	10, NO_2	10, 1.50 mol O_2
1.00 mol	0.00 mol	12, 2.00 mol	12, O_2	12, 2.00 mol NO
0.200 mol	0.200 mol	16, 0.200 mol	17, NO_2	16, 0.400 mol O_2
33.00 g	46.01 g	19, 92.02 g	20, none	21, none
64.00 g	92.02 g	22, 46.01 g	22, O_2	24, 30.01 g NO
64.00 g	30.01 g	25, 28.76 g	26, O_2	27, 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

16.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.5125 mol O_2

20.00 g NO_2 = 1 mol NO_2 /30.01 g NO_2 = 0.6664 mol NO_2

actual mole ratio = 0.6664 mol NO_2 /0.5125 mol O_2 = 2.132 mol NO_2 /1.000 mol O_2

Because the actual mole ratio of NO_2 / O_2 is larger than the balanced equation mole ratio of NO_2 / O_2 , there is an excess of NO_2 ; O_2 is the limiting reactant.

Mass of NO used = 0.5125 mol O_2 = 2 mol NO /1 mol O_2 = 0.4250 mol NO

0.4250 mol NO = 30.01 g NO /1 mol NO = 12.76 g NO

Mass of NO_2 produced = 0.4250 mol NO_2 = 46.01 g NO_2 /1 mol NO_2 = 20.76 g NO_2

Excess NO_2 = 20.00 g NO_2 = 18.76 g NO_2 = 1.24 g NO_2

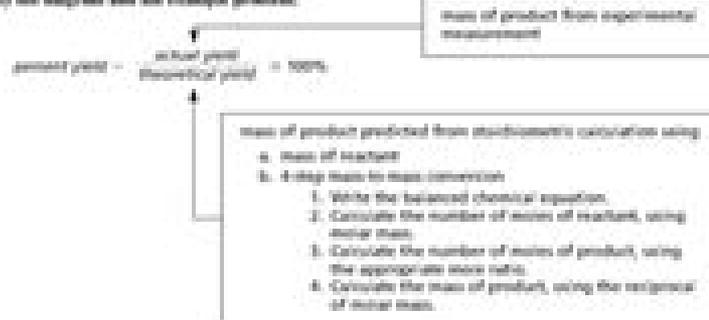
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CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.4 Percent Yield

In your textbook, read about the percent of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of yellow oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga , correct
- number of moles of reactant: 7.00 g Ga = $\frac{69.72 \text{ g } Ga}{1 \text{ mol } Ga}$, 7.00 g Ga = 1 mol Ga /69.72 g Ga
- number of moles of product: 0.100 mol Ga = $\frac{2 \text{ mol } Ga_2O_3}{1 \text{ mol } Ga}$, 0.100 mol Ga = 2 mol Ga_2O_3 /4 mol Ga
- theoretical yield: 0.100 mol Ga_2O_3 = $\frac{187.44 \text{ g } Ga_2O_3}{1 \text{ mol } Ga_2O_3}$, correct
- percent yield: $\frac{7.42 \text{ g } Ga_2O_3}{7.42 \text{ g } Ga_2O_3} \times 100$, 7.42 g Ga_2O_3 /7.42 g Ga_2O_3 = 100

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1. Understanding the eBook Standardized Test Practice Chapter 12 Chemistry Answers
 - The Rise of Digital Reading Standardized Test Practice Chapter 12 Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Standardized Test Practice Chapter 12 Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Standardized Test Practice Chapter 12 Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Standardized Test Practice Chapter 12 Chemistry Answers
 - Personalized Recommendations
 - Standardized Test Practice Chapter 12 Chemistry Answers User Reviews and Ratings
 - Standardized Test Practice Chapter 12 Chemistry Answers and Bestseller Lists

5. Accessing Standardized Test Practice Chapter 12 Chemistry Answers Free and Paid eBooks
 - Standardized Test Practice Chapter 12 Chemistry Answers Public Domain eBooks
 - Standardized Test Practice Chapter 12 Chemistry Answers eBook Subscription Services
 - Standardized Test Practice Chapter 12 Chemistry Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Standardized Test Practice Chapter 12 Chemistry Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Standardized Test Practice Chapter 12 Chemistry Answers
 - Interactive Elements Standardized Test Practice Chapter 12 Chemistry Answers
8. Staying Engaged with Standardized Test Practice Chapter 12 Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Standardized Test Practice Chapter 12 Chemistry Answers
9. Balancing eBooks and Physical Books Standardized Test Practice Chapter 12 Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Standardized Test Practice Chapter 12 Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Standardized Test Practice Chapter 12 Chemistry Answers
 - Setting Reading Goals Standardized Test Practice Chapter 12 Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Standardized Test Practice Chapter 12 Chemistry Answers
 - Fact-Checking eBook Content of Standardized Test Practice Chapter 12 Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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