

Honors Chemistry Midterm: Review Packet

PART 1

1. Write in the number of significant figures for each of the following measurements:

3 3.00 mL 3 0.00290 g 4 50.00 m 2 0.070 kg 1 400 L

2. Using the following masses: 0.0562 kg, 124.213 g and 1635 mg. *← convert to grams*
 The answer should be reported as: g

56.2g 124.213g 1.635g

3. a) A 5.75 mL sample of mercury has a measured mass of 77.05 g. The density is 13.49/mL

$$D = \frac{m}{V} = \frac{77.05g}{5.75mL} =$$

b) Mercury's accepted density is 13.53 g/mL. The % error in your measurement is: .96170

$$\frac{|actual - exp|}{actual} \times 100 \quad \frac{|13.53g/mL - 13.49g/mL|}{13.53g/mL} \times 100$$

4. In the diagram below:



Is this person accurate? no - not in center

Is this person precise? yes - close together

5. Identify the following as a chemical change (CC), physical change (PC), chemical property (CP), or physical property (PP):

a) burning wood CC b) mass of wood PP

c) flammability of wood CP d) chopping up wood PC

Honors Chemistry Semester Review Packet Answers

Mary Millhollon, Richard H. Langley



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