

3. Atoms, elements and compounds



Atomic mass = 23

Proton number = 11

Neutron number = 12 (23 - 11)

mass of electron = $\frac{1}{1836}$

ISOTOPE: atoms of the same element with the same proton number but a different neutron number

- Have the same properties due to the same number of electrons on the outer shell

Radioisotope

- testing for leaks in pipes
- treating cancer

IONIC BONDING

Bonding: oppositely charged ions are strongly attracted by electrostatic force + involves electron transfer

metal + non metal

- metals give electrons to non-metals

Properties: strong bonds due to attractive forces

- conduct electricity
- high mp/bp
- charged particles
- soluble in water

Macromolecular compounds

SILICON (IV) OXIDE

Properties



COVALENT BONDING

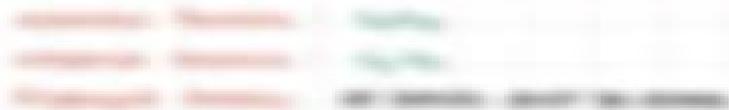
Bonding: nuclei of atoms and the shared pair of electrons are strongly attracted by electrostatic forces

non metal + non metal

Properties: weak bonds = low mp/bp

- insoluble
- don't conduct

3. Bonding



... the strong ... of

A large box containing a list of properties and examples of covalent bonding. The text is partially illegible but appears to list various characteristics and examples of covalent compounds.

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RJ Alexander



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