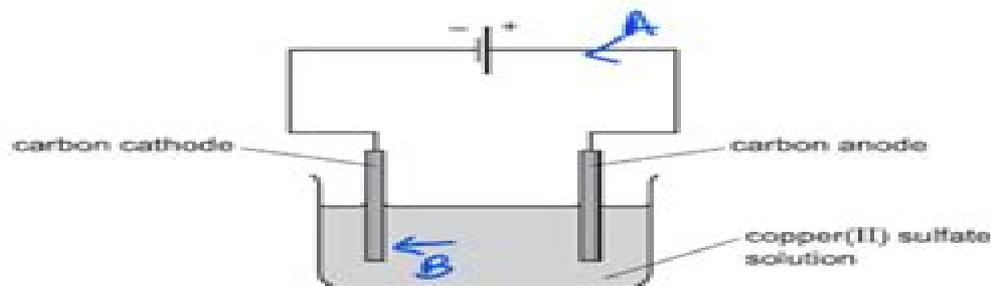


- 4 Copper(II) sulfate solution was electrolysed using the apparatus shown.



- (a) (i) Draw an arrow on the diagram to show the direction of movement of electrons in the wire. Label the arrow **A**.  
*arrow labelled A on or near wire going in an anti-clockwise direction* [1]
- (ii) Draw an arrow on the diagram to show the direction of movement of positive ions in the copper(II) sulfate solution. Label the arrow **B**.  
*arrow labelled B in electrolyte pointing towards the cathode* [1]
- (b) Oxygen was formed at the anode and copper was formed at the cathode.
- (i) The ionic half-equation for the formation of oxygen is shown.



Explain why this reaction is oxidation.

*electrons are lost*

- ..... [1]
- (ii) Write the ionic half-equation for the formation of copper at the cathode.  
*M1 Cu<sup>2+</sup> ions on left*  
 .....  
*M2 rest of equation correct and correctly balanced*  
 (Cu<sup>2+</sup> + 2e<sup>-</sup> → Cu)

- (c) The electrolysis was repeated using copper electrodes in place of carbon electrodes.

State and explain what happens to the masses of the anode and the cathode during this electrolysis.

*M1 anode mass decreases*

.....  
*M2 copper lost as ions OR copper (atoms) becomes ions OR*



.....  
*M3 cathode mass increases*

.....  
*M4 copper deposited / layer of copper forms / copper collected at cathode*



..... [4]

[Total: 9]

# Igcse Electrolysis Past Papers

**Arshad Iqbal**



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