

Chapter 1

Problem Solutions

1.1

- (a) fcc: 8 corner atoms $\times 1/8 = 1$ atom
6 face atoms $\times 1/2 = 3$ atoms
Total of 4 atoms per unit cell

- (b) bcc: 8 corner atoms $\times 1/8 = 1$ atom
1 enclosed atom = 1 atom
Total of 2 atoms per unit cell

- (c) Diamond: 8 corner atoms $\times 1/8 = 1$ atom
6 face atoms $\times 1/2 = 3$ atoms
4 enclosed atoms = 4 atoms
Total of 8 atoms per unit cell

1.2

- (a) 4 Ga atoms per unit cell

$$\text{Density} = \frac{4}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ga} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

4 As atoms per unit cell, so that

$$\underline{\text{Density of As} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

- (b)

8 Ge atoms per unit cell

$$\text{Density} = \frac{8}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ge} = 4.44 \times 10^{22} \text{ cm}^{-3}}$$

1.3

- (a) Simple cubic lattice; $a = 2r$

$$\text{Unit cell vol} = a^3 = (2r)^3 = 8r^3$$

$$1 \text{ atom per cell, so atom vol.} = (1) \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{\left(\frac{4\pi r^3}{3} \right)}{8r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 52.4\%}$$

- (b) Face-centered cubic lattice

$$d = 4r = a\sqrt{2} \Rightarrow a = \frac{d}{\sqrt{2}} = 2\sqrt{2}r$$

$$\text{Unit cell vol} = a^3 = (2\sqrt{2}r)^3 = 16\sqrt{2}r^3$$

$$4 \text{ atoms per cell, so atom vol.} = 4 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{4 \left(\frac{4\pi r^3}{3} \right)}{16\sqrt{2}r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 74\%}$$

- (c) Body-centered cubic lattice

$$d = 4r = a\sqrt{3} \Rightarrow a = \frac{4}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{4}{\sqrt{3}}r \right)^3$$

$$2 \text{ atoms per cell, so atom vol.} = 2 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{2 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{4}{\sqrt{3}}r \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 68\%}$$

- (d) Diamond lattice

$$\text{Body diagonal} = d = 8r = a\sqrt{3} \Rightarrow a = \frac{8}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{8r}{\sqrt{3}} \right)^3$$

$$8 \text{ atoms per cell, so atom vol.} = 8 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{8 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{8r}{\sqrt{3}} \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 34\%}$$

1.4

From Problem 1.3, percent volume of fcc atoms is 74%; Therefore after coffee is ground,

$$\text{Volume} = 0.74 \text{ cm}^3$$

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web shoplifting is a practice that has been engaged in for centuries but it was only after the civil war that the prevalence of shoplifting and societal awareness of it became significant in the 1860s the typical shoplifter was from the lower classes by 1900 it was an upper class woman who