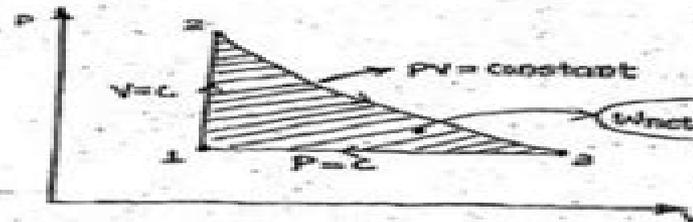


Solution (Assignment 1)

Given:

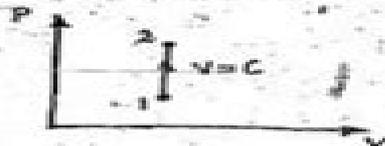


$T_1 = 350 \text{ K}$
 $T_2 = 3T_1 = 3(350) = 1050 \text{ K}$
 $T_3 = T_2$ (isothermal process)
 $T_3 = 1050 \text{ K}$
 $n = 1$

$V_2 = V_1$
 $T_3 = T_2$

Net work done / kg of gas = $w_{12} + w_{23} + w_{31}$

For process 1-2



constant volume process

$$\frac{P_2}{P_1} = \frac{T_2}{T_1} = \frac{1050}{350}$$

$$\frac{P_2}{P_1} = 3$$

$$P_2 = 3P_1$$

$$w_{12} = 0$$



isothermal process

$$PV = \text{constant}$$

$$w_{23} = P_2 V_2 \ln \frac{V_3}{V_2}$$

$$= mRT_2 \ln \frac{V_3}{V_2}$$

$$= (1)(R)1050 \ln \frac{V_3}{V_2}$$

$$= R(1050) \ln 3$$

$$= 1050 \times 1.0986 \times R$$

$$w_{23} = 1153.542909 R$$



constant pressure process

$$P = C$$

$$\frac{V_3}{V_1} = \frac{T_3}{T_1}$$

$$\frac{V_3}{V_1} = \frac{T_3}{T_1} = \frac{1050}{350}$$

$$\frac{V_3}{V_1} = 3$$

$$w_{31} = P(V_1 - V_3)$$

$$= P V_1 - P V_3$$

$$= mRT_1 - mRT_3$$

$$= RT_1 - RT_3$$

$$= R(T_1 - T_3)$$

$$= R(350 - 1050)$$

$$w_{31} = -700 R$$

Net work done / kg of gas = $w_{12} + w_{23} + w_{31}$

$$= 0 + 1153.5429 R - 700 R$$

$$\frac{w_{\text{net}}}{\text{kg of gas}} = 453.5429 R \text{ Ans}$$

Thermodynamics Mechanical Engineering Notes

AW Chickering



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