

AP Chemistry Unit 3 Practice Test
Solids, Liquids, Gases and Solutions

Name _____
Date _____ Block _____

1. A real gas would act most ideal at
(A) 1 atm and 273 K (D) 0.5 atm and 546 K
(B) 10 atm and 547 K (E) 0.5 atm and 273 K
(C) 10 atm and 273 K
2. Zinc reacts with aqueous sulfuric acid to form hydrogen gas:
$$\text{Zn (s)} + \text{H}_2\text{SO}_4 \text{ (aq)} \rightarrow \text{ZnSO}_4 \text{ (aq)} + \text{H}_2 \text{ (g)}$$

In an experiment, 201 mL of wet H_2 is collected over water at 27°C and a barometric pressure of 733 torr. The vapor pressure of water at 27°C is 26.74 torr. The partial pressure of hydrogen in this experiment is _____ atm.
(A) 0.929 (D) 1.00
(B) 760 (E) 0.964
(C) 706
3. A mixture of gases contains 1.5 moles of oxygen, 3.0 moles of nitrogen, and 0.5 mole of water vapor. If the total pressure is 700 mmHg, what is the partial pressure of the nitrogen gas?
(A) 70 mmHg (D) 350 mmHg
(B) 210 mmHg (E) 420 mmHg
(C) 280 mmHg
4. Equal numbers of moles of $\text{CO}_2(\text{g})$, $\text{N}_2(\text{g})$, and $\text{NH}_3(\text{g})$ are placed in a sealed vessel at room temperature. If the vessel has a pinhole-size leak, which of the following will be true after some of the gas mixture has effused?
I. The mole fraction of CO_2 in the sample will increase.
II. The N_2 will effuse the fastest since it is the lightest.
III. All gases will effuse at the same rate since the temperature is held constant.
(A) I only (C) I and II (E) I, II, and III
(B) III only (D) II and III
5. A sample of oxygen gas was found to effuse at a rate equal to two times that of an unknown gas. The molecular weight of the unknown gas is _____ g/mol.
(A) 128
(B) 8
(C) 16
(D) 64

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**Philip E. Dumas, Ronald M. Fikar, Kevin
R. Reel, Jay M. Templin**

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Table of Contents Ap Chemistry Chapter 3 Practice Test Pluteo

1. Understanding the eBook Ap Chemistry Chapter 3 Practice Test Pluteo
 - The Rise of Digital Reading Ap Chemistry Chapter 3 Practice Test Pluteo
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Chapter 3 Practice Test Pluteo
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chemistry Chapter 3 Practice Test Pluteo
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Chapter 3 Practice Test Pluteo
 - Personalized Recommendations

- Ap Chemistry Chapter 3 Practice Test Pluteo User Reviews and Ratings
- Ap Chemistry Chapter 3 Practice Test Pluteo and Bestseller Lists
- 5. Accessing Ap Chemistry Chapter 3 Practice Test Pluteo Free and Paid eBooks
 - Ap Chemistry Chapter 3 Practice Test Pluteo Public Domain eBooks
 - Ap Chemistry Chapter 3 Practice Test Pluteo eBook Subscription Services
 - Ap Chemistry Chapter 3 Practice Test Pluteo Budget-Friendly Options
- 6. Navigating Ap Chemistry Chapter 3 Practice Test Pluteo eBook Formats
 - ePub, PDF, MOBI, and More
 - Ap Chemistry Chapter 3 Practice Test Pluteo Compatibility with Devices
 - Ap Chemistry Chapter 3 Practice Test Pluteo Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ap Chemistry Chapter 3 Practice Test Pluteo
 - Highlighting and Note-Taking Ap Chemistry Chapter 3 Practice Test Pluteo
 - Interactive Elements Ap Chemistry Chapter 3 Practice Test Pluteo
- 8. Staying Engaged with Ap Chemistry Chapter 3 Practice Test Pluteo
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Chemistry Chapter 3 Practice Test Pluteo
- 9. Balancing eBooks and Physical Books Ap Chemistry Chapter 3 Practice Test Pluteo
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chemistry Chapter 3 Practice Test Pluteo
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ap Chemistry Chapter 3 Practice Test Pluteo
 - Setting Reading Goals Ap Chemistry Chapter 3 Practice Test Pluteo
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ap Chemistry Chapter 3 Practice Test Pluteo
 - Fact-Checking eBook Content of Ap Chemistry Chapter 3 Practice Test Pluteo

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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