

Chapter 17: Thermochemistry*Section 1: The Flow of Energy – Heat and Work*

1. What is thermochemistry? **The study of the energy changes that occur in changes of state or during chemical reactions.**
2. Heat always flows from a **warmer** object to a **cooler** object.
3. Given that 1 cal = 4.184 J:
 - a. How many calories are in 325.9 J? **77.89 cal**
 - b. How many J are in 34.7 cal? **145.18 J**
4. What is the Law of Conservation of Energy (aka the First Law of Thermodynamics)? **States that in chemical and physical processes, energy is neither created nor destroyed – it is transferred or conserved.**
5. In an exothermic reaction:
 - a. Heat is **released** / absorbed
 - b. q is **positive** / **negative**
 - c. ΔH is **positive** / **negative**
 - d. Heat is a **reactant** / **product**
6. In an endothermic reaction:
 - a. Heat is **released** / **absorbed**
 - b. q is **positive** / **negative**
 - c. ΔH is **positive** / **negative**
 - d. Heat is a **reactant** / **product**
7. Diagram **A** represents an exothermic reaction, whereas diagram **B** represents an endothermic reaction:
8. What is the difference between heat capacity and specific heat capacity? **Specific heat is the amount of energy to increase the temperature of 1g of substance by 1°C**
9. The temperature of a 34.1 g sample of a substance increases from 29 C to 51 C when the sample absorbs 912 J of heat. What is the specific heat of the substance? **1.126 J/g°C**
10. When 867 J of heat are added to 9.3 g of a substance at STP, the temperature increases from 0 C to 14 C. What is the specific heat of the substance? **6.66 J/g°C**

Chapter 17 Thermochemistry Section Review Answers

SA Dillow



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