

# Specific Heat

Chem 101/102 Lab 10-1

Name: \_\_\_\_\_

Use the data table on this table for answers. Use 3 significant figures.

Substance	Specific Heat Capacity (J/g °C)
water	4.184
aluminum	0.897
iron	0.449
copper	0.385
gold	0.129
lead	0.128

Metal Properties	
Aluminum	10.000 g
Iron	10.000 g
Copper	10.000 g
Gold	10.000 g
Lead	10.000 g

1. A 10.000 g metal sample is heated to 100.0 °C by boiling. It is then immersed in 20.0 g of water at 20.0 °C. The final temperature of the water is 25.0 °C. What is the specific heat capacity of the metal? **0.897 J/g °C**
2. An unknown metal (mass = 10.0 g) is heated to 100.0 °C, is submerged in 20.0 g of water at 20.0 °C. The final temperature of the water is 25.0 °C. What is the specific heat capacity of the metal? **0.897 J/g °C**
3. Determine the temperature change that will occur when 10.0 g of water is supplied 100 g of gold. **0.003 °C**
4. A 10.0 g metal sample is supplied 100 J of energy. The temperature of the metal increases by 10.0 °C. Determine the mass of the metal sample. **10.00 g**
5. An unknown metal sample of 10.0 g is heated to 100.0 °C. It is submerged in 20.0 g of water at 20.0 °C. The final temperature of the water is 25.0 °C. Determine the specific heat of the metal. **0.897 J/g °C**
6. A 10.0 g metal sample is submerged in 20.0 g of water at 20.0 °C. The temperature of the water increases by 10.0 °C. Determine the specific heat of the metal. **0.897 J/g °C**
7. What is the mass of a metal sample that is heated to 100.0 °C, placed in 20.0 g of water at 20.0 °C, and the final temperature of the water is 25.0 °C? **10.00 g**
8. How much energy is required to heat 10.0 g of water from 20.0 °C to 100.0 °C? **3347 J**
9. A 10.0 g metal sample is heated to 100.0 °C. It is submerged in 20.0 g of water at 20.0 °C. The final temperature of the water is 25.0 °C. Determine the specific heat of the metal. **0.897 J/g °C**
10. A 10.0 g metal sample is heated to 100.0 °C. It is submerged in 20.0 g of water at 20.0 °C. The final temperature of the water is 25.0 °C. Determine the specific heat of the metal. **0.897 J/g °C**

**0.897 J/g °C**      **0.897 J/g °C**

# Specific Heat Chem Worksheet 16 1 Answer Key

**AW Rasmussen**



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