

This is the required relation between r and R .

15. Copper crystallises into a fcc lattice with edge length 3.61×10^{-8} cm. Show that the calculated density is in agreement with its measured value of 8.92 g cm^{-3} .

15. Edge length, $a = 3.61 \times 10^{-8}$ cm

As the lattice is fcc type, the number of atoms per unit cell, $z = 4$

Atomic mass, $M = 63.5 \text{ g mol}^{-1}$

We also know that, $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$ Applying the relation:

$$\begin{aligned}d &= \frac{zM}{a^3 N_A} \\ &= \frac{4 \times 63.5 \text{ g mol}^{-1}}{(3.61 \times 10^{-8} \text{ cm})^3 \times 6.022 \times 10^{23} \text{ mol}^{-1}} \\ &= 8.97 \text{ g cm}^{-3}\end{aligned}$$

The measured value of density is given as 8.92 g cm^{-3} . Hence, the calculated density 8.97 g cm^{-3} is in agreement with its measured value.

16. Analysis shows that nickel oxide has the formula $\text{Ni}_{0.98}\text{O}_{1.00}$. What fractions of nickel exist as Ni^{2+} and Ni^{3+} ions?

16. Given formula of nickel oxide = $\text{Ni}_{0.98}\text{O}_{1.00}$

Therefore, Ni : O = 0.98:1.00

Total charge on O^{2-} ion = $100 \times (-2) = -200$

Now, let the number of Ni^{2+} ions = x

Therefore, number of Ni^{3+} = $98 - x$

Now, since compound is neutral

Therefore, no. of Ni^{2+} ions + no. of Ni^{3+} ions + no. of O^{2-} ions = 0

$$x(-2) + (98 - x) \times (-3) + (-200) = 0$$

$$-2x - 294 + 3x - 200 = 0$$

$$x - 94 = 0$$

$$x = 94$$

Therefore, the number of Ni^{2+} ions = 94

Number of Ni^{3+} ions = $98 - 94 = 4$

$$\text{Now, fraction of nickel exists as } \text{Ni}^{2+} = \frac{94}{98} = 0.959$$

$$\text{Fraction of nickel exists as } \text{Ni}^{3+} \text{ ions} = \frac{4}{98} = 0.041$$

17. What is a semiconductor? Describe the two main types of semiconductors and contrast their conduction mechanism.

Solid State Physics Chapter 1 Solutions

David Frank Eggers (Jr.)



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