

Standard Thermodynamic Values at 25°C

Please note that enthalpy and free energy values are given in kJ/mol while entropy values are given in J/mol K.

Formula	State	ΔH_f°	S°	ΔG_f°
(BOC) ₂	(g)	-1633.43	380.74	-1550.17
(CN) ₂	(g) - cyanogen	308.05	242.25	297.19
(NH ₂) ₂ CO	(s) - urea	-333.51	104.60	-196.83
(NH ₂) ₂ O	(s)	-490.70	267.52	-267.11
(NH ₂) ₂ SO ₄	(s) hexagonal	-2683.69	280.24	-2365.55
(NH ₂) ₂ SO ₃	(s)	-1180.85	220.08	-901.90
[Co(NH ₂) ₂ CO] ₂	(s)	-1017.13	366.10	-582.83
[Co(NH ₂) ₂ NO ₂] ₂	(s)	-2088.68	347.27	-418.40
[Co(NH ₂) ₂ NO ₂] ²⁺	(aq)	-633.37	179.91	-172.80
[Co(NH ₂) ₂ CO] ₂	(s)	-1034.70	635.05	-221.75
[Co(NH ₂) ₂ NO ₂] ₂	(s)	-1283.98	447.69	-536.72
[Co(NH ₂) ₂ Br] ₂	(s)	-1002.90	325.10	-505.24
Ac	(g)	390.30	0.00	0.00
Ac	(l)	0.00	62.76	0.00
Ac ₂ O ₃	(s)	-1807.70	0.00	-1778.20
AcBr ₃	(s)	-920.88	0.00	0.00
AcCl ₃	(s)	-1133.86	0.00	0.00
AcF ₃	(s)	-1995.77	0.00	0.00
AcOF	(s)	-1108.76	0.00	0.00
Ag	(s)	0.00	42.55	0.00
Ag	(l)	284.55	172.89	245.68
Ag ⁺	(aq)	105.58	72.68	77.12
Ag ₂	(g)	409.09	257.02	358.78
Ag ₂ C ₂ O ₄	(s)	-673.21	209.20	-584.09
Ag ₂ CO ₃	(s)	-505.85	167.16	-436.81
Ag ₂ O ⁺	(s)	-733.74	217.57	-641.83
Ag ₂ MnO ₄	(s)	-840.57	213.38	-748.10
Ag ₂ O	(s)	-31.05	123.34	-11.21
Ag ₂ O ₂	(s)	-34.27	137.15	27.61

Formula	State	ΔH_f°	S°	ΔG_f°
Ag ₂ O ₂	(s)	33.89	100.42	121.34
Ag ₂ S	(s) alpha orthorhombic	-32.59	144.05	-40.67
Ag ₂ S	(s) beta	-29.41	150.62	-39.86
Ag ₂ Se	(s)	-37.66	150.71	-44.35
Ag ₂ SeO ₃	(s)	-365.26	230.12	-304.18
Ag ₂ SeO ₄	(s)	-420.49	248.53	-334.30
Ag ₂ SO ₃	(s)	-490.78	158.16	-411.29
Ag ₂ SO ₄	(s)	-715.88	200.41	-618.48
Ag ₂ Te	(s)	-37.24	154.81	-43.10
AgBr	(s)	-100.37	107.11	-96.90
AgBrO ₃	(s)	-37.20	152.72	-54.39
AgCl	(s)	-127.07	96.23	-109.80
AgClO ₂	(s)	8.79	134.56	73.73
AgClO ₃	(s)	-80.3	142.0	-64.5
AgClO ₄	(s)	-31.1		
AgCN	(s)	146.02	107.19	156.90
AgF+2H ₂ O	(s)	-800.82	174.89	-671.11
AgI	(s)	-61.84	115.48	-66.19
AgIO ₃	(s)	-171.53	149.37	-93.72
AgI ₃	(s)	308.78	104.18	176.14
AgNO ₂	(s)	-45.06	128.20	19.08
AgNO ₃	(s)	-124.39	140.92	-33.47
AgO	(s)	-11.42	57.78	14.23
AgOCN	(s)	-95.40	121.34	-58.16
AgReO ₄	(s)	-716.38	153.13	-635.55
AgSCN	(s)	87.86	130.96	101.38
Al	(s)	0.00	28.33	0.00
Al	(l)	8.66	35.23	6.61
Al	(g)	326.35	164.43	285.77

Standard Thermodynamic Values At 25 C Chemistry Reference

William M. Haynes



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Reagents Major update of Density of Solvents as a Function of Temperature Major update of Dependence of Boiling Point on Pressure Section 16 Health and Safety Information Major update of Threshold Limits for Airborne Contaminants Appendix A Major update of Mathematical Tables Appendix B Update of Sources of Physical and Chemical Data *Standard Aqueous Electrode Potentials and Temperature Coefficients at 250C*, 1964 *Lange's Handbook of Chemistry, 70th Anniversary Edition* James Speight, 2005-01-10 A standard reference for chemists for 70 years this new Sixteenth Edition features an enormous compilation of facts data tabular material and experimental findings in every area of chemistry Included in this massive compendium are listings of the properties of approximately 4 400 organic and 1 400 inorganic compounds This Sixteenth Edition offers 40% new or extensively revised content and starting with this edition the author includes equations that allow users to calculate important values such as temperature and pressure Contents Organic Compounds General Information Conversion Tables and Mathematics Inorganic Compounds Properties of Atom Radicals and Bonds Physical Properties Thermodynamic Properties Spectroscopy Electrolytes Electromotive Force and Chemicals Physicochemical Relationships Polymers Rubbers Fats Oils and Waxes Practical Laboratory Information **Lange's Handbook of Chemistry** John Aurie Dean, Norbert Adolph Lange, 1985 **Physical Chemistry** Ignacio Tinoco (Jr.), Kenneth Sauer, James C. Wang, 1985 **Physical Chemistry** Ignacio Tinoco, 2002 This best selling volume presents the principles and applications of physical chemistry as they are used to solve problems in biology and medicine The First Law the Second Law free energy and chemical equilibria free energy and physical Equilibria molecular motion and transport properties kinetics rates of chemical reactions enzyme kinetics the theory and spectroscopy of molecular structures and interactions molecular distributions and statistical thermodynamics and macromolecular structure and X ray diffraction For anyone interested in physical chemistry as it relates to problems in biology and medicine **Physical Chemistry** Ira N. Levine, 2009 Ira N Levine's sixth edition of Physical Chemistry provides students with an in depth fundamental treatment of physical chemistry At the same time the treatment is made easy to follow by giving full step by step derivations clear explanations and by avoiding advanced mathematics unfamiliar to students Necessary math and physics have thorough review sections Worked examples are followed by a practice exercise National Standard Reference Data Series United States. National Bureau of Standards, 1964 **Gmelin Handbook of Inorganic Chemistry**, 1987 **Chemistry and Industry**, 1992 **Government Reports Announcements & Index**, 1996 **Government Reports Announcements & Index**, 1988 **Industrial Arts Index**, 1930 **U.S. Government Research and Development Reports**, 1965 **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 **The Industrial Arts Index**, 1930 *Russian Journal of Inorganic Chemistry*, 1988

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