

Stoichiometry Lab

Single Replacement Reactions: Aluminum and Copper

Pre-Lab Questions

1. Define Stoichiometry
2. What is a mole-to-mole ratio and why is it important in stoichiometry? Find an example of a balanced equation (not given to you) and use to explain.
3. Write the complete, balanced equation showing what happens when aluminum foil is placed into copper (II) chloride solution.
 - a. Write two possible mole-to-mole ratios and explain when you would use them.
 - b. What type of reaction is this? Will it happen and how do you know?
4. How do you find theoretical yield?
5. How do you find actual yield?
6. What is the formula for % yield?

Materials

- | | |
|----------------------------|---------------------------------|
| 1. One 50 or 100 mL beaker | 8. 25 mL of D.I. water |
| 2. One glass stir rod | 9. One ring stand |
| 3. One scoopula | 10. One piece of filter paper |
| 4. One digital balance | 11. One funnel |
| 5. Two weigh boats | 12. One 150 mL Erlenmeyer flask |
| 6. 0.2 g of aluminum foil | |
| 7. 2 g of CuCl_2 | |

Procedure

Watch the video to see the experiment and to get the data you will record in your data table. <https://youtu.be/2a8-VNxtTic> Be sure to also record DETAILED qualitative and quantitative observations for the aluminum and copper (II) chloride solution in the data table.

Day 1 (if we were doing this in class)

1. Place a weigh boat on the balance and tare the instrument.
2. Weigh out 0.2 g of aluminum foil.
3. Tear the aluminum foil into smaller pieces to increase the reaction speed. Keep in weigh boat until needed.
4. Repeat step 1 with your other weigh boat. Using the scoopula, weigh out 2 g of copper (II) chloride into the weigh boat.
5. Pour the copper (II) chloride crystals into your beaker and add your 25 mL of D.I. water. Stir with a glass rod until fully dissolved.
6. Record all observations of your reactants before the reaction takes place.

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RM Cervero



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