

### Concentration

Mass

Volume

### Concentration of solution

#### Quantities

- 1. Mass concentration of solution
- 2. Molar
- 3. Molar

#### Mass

How do you calculate the mass of a substance in a solution with a certain volume and concentration? (The amount of mass in a certain volume of solution of certain concentration)

#### Learning objectives

After completing this lesson, you should be able to:

- 1. calculate the concentration of a solution in g dm<sup>-3</sup> of a given volume, given its mass (concentration in g dm<sup>-3</sup>)
- 2. calculate the mass of a substance in a given volume of a solution, given its concentration in g dm<sup>-3</sup>

#### Working the volume

Example 1: How many grams of sodium hydroxide are there in 250 cm<sup>3</sup> of a solution with a concentration of 40 g dm<sup>-3</sup>?

Solution: We know the concentration of the solution is 40 g dm<sup>-3</sup>. This means that in 1 dm<sup>3</sup> of the solution, there are 40 g of sodium hydroxide.

$$\text{Concentration (g dm}^{-3}\text{)} = \frac{\text{Mass of substance (g)}}{\text{Volume (dm}^3\text{)}}$$

Therefore, the mass of sodium hydroxide in 250 cm<sup>3</sup> of a solution with a concentration of 40 g dm<sup>-3</sup> is:

$$\text{Concentration (g dm}^{-3}\text{)} = \frac{\text{Mass of substance (g)}}{\text{Volume (dm}^3\text{)}}$$

Example 2: How many dm<sup>3</sup> of a solution with a concentration of 20 g dm<sup>-3</sup> are there in 400 g of sodium hydroxide? (The volume of a solution with a certain mass and concentration)

# Ssd Solution Formula

**DJ Losen**



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