

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{FeSCNH}^{3+}(\text{aq})$$
4. In part 4, what is the purpose of adding potassium thiocyanate to 100 mL of 0.1 M FeCl₃ solution?
To observe the effect of the central group and observe the effect of the ligand. The purpose of adding FeCl₃ to the solution is to observe the effect of the central group and observe the effect of the ligand.
5. What is the purpose of adding distilled water to part 4?
To observe the effect of the central group and observe the effect of the ligand.
6. What is the purpose of adding 10 mL of 0.1 M FeCl₃ to part 4?
To observe the effect of the central group and observe the effect of the ligand.
7. Write the equation for the reaction between Fe³⁺ and OH⁻ ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$
8. Write the equation for the reaction between Fe³⁺ and OH⁻ ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$
9. What is the purpose of adding 10 mL of 0.1 M FeCl₃ to part 4?
To observe the effect of the central group and observe the effect of the ligand.

Investigating Chemical Equilibrium Lab 12a Answer

Emilie Sanchez



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