



## Student Exploration: Discharge Rate

(My answers are indicated in red.)

**Problem:** Suppose a certain chemical reaction occurs with constant temperature, pressure, and volume. Then the rate of change of the concentration of the reactants is proportional to the concentration of the reactants.

**What knowledge, questions, or ideas do you bring to this problem?**

1. A differential equation of the form  $y' = ky$  (where  $k$  is a constant) has the solution  $y = Ce^{kt}$ . In this case,  $y$  is the concentration of the reactants,  $t$  is time, and  $C$  is a constant.
2. If the concentration of the reactants is  $y$  at time  $t$ , then the rate of change of the concentration is  $y'$ . The differential equation is  $y' = ky$ .

**What do you know?**

I know that the differential equation is  $y' = ky$ . I also know that the solution is  $y = Ce^{kt}$ .

**What do you want to know?** I want to know how to find the concentration of the reactants at a given time  $t$ . I also want to know how to find the rate of change of the concentration of the reactants at a given time  $t$ .

**Strategy:** I will use the differential equation to find the concentration of the reactants.

$$y' = ky \quad \text{for } y(0) = 1000.$$

1. What is the concentration of the reactants at time  $t$ ? (I will use the differential equation to find the concentration of the reactants at time  $t$ .)

Time $t$	Concentration $y$	Rate of Change $y'$
0	1000	0
1	1000e <sup>k</sup>	1000ke <sup>k</sup>
2	1000e <sup>2k</sup>	2000ke <sup>2k</sup>
3	1000e <sup>3k</sup>	3000ke <sup>3k</sup>

**Answers:**  $y = 1000e^{kt}$ ,  $y' = 1000ke^{kt}$   
**Questions:** How do I find  $k$ ? How do I find  $C$ ?

**What do you think?** Is the reaction first-order? Yes, the equation is  $y' = ky$ .

**What do you know?**

$$y' = ky \quad \text{for } y(0) = 1000.$$

**What do you want to know?**

I want to know how to find the concentration of the reactants at a given time  $t$ .

**Strategy:**



# Explore Learning Stoichiometry Answers

**Rachel Sandford**



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### **Table of Contents Explore Learning Stoichiometry Answers**

1. Understanding the eBook Explore Learning Stoichiometry Answers
  - The Rise of Digital Reading Explore Learning Stoichiometry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Stoichiometry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Explore Learning Stoichiometry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Stoichiometry Answers

- Personalized Recommendations
  - Explore Learning Stoichiometry Answers User Reviews and Ratings
  - Explore Learning Stoichiometry Answers and Bestseller Lists
5. Accessing Explore Learning Stoichiometry Answers Free and Paid eBooks
    - Explore Learning Stoichiometry Answers Public Domain eBooks
    - Explore Learning Stoichiometry Answers eBook Subscription Services
    - Explore Learning Stoichiometry Answers Budget-Friendly Options
  6. Navigating Explore Learning Stoichiometry Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Explore Learning Stoichiometry Answers Compatibility with Devices
    - Explore Learning Stoichiometry Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Explore Learning Stoichiometry Answers
    - Highlighting and Note-Taking Explore Learning Stoichiometry Answers
    - Interactive Elements Explore Learning Stoichiometry Answers
  8. Staying Engaged with Explore Learning Stoichiometry Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Explore Learning Stoichiometry Answers
  9. Balancing eBooks and Physical Books Explore Learning Stoichiometry Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Explore Learning Stoichiometry Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Explore Learning Stoichiometry Answers
    - Setting Reading Goals Explore Learning Stoichiometry Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Explore Learning Stoichiometry Answers

- Fact-Checking eBook Content of Explore Learning Stoichiometry Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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