

## Lewis Structure Worksheet #2

Key

Draw Lewis Structures for each of the following molecules or ions:

1. Br <sub>2</sub> 14 $\text{:}\ddot{\text{Br}}-\ddot{\text{Br}}\text{:}$	2. BH <sub>3</sub> 6 $\begin{array}{c} \text{H} \\   \\ \text{B} \\   \\ \text{H} \end{array}$
3. O <sub>2</sub> 12 $\text{:}\ddot{\text{O}}=\ddot{\text{O}}\text{:}$	4. PS <sub>3</sub> <sup>-</sup> 24 $\begin{array}{c} \text{:}\ddot{\text{S}} \\   \\ \text{P} \\   \\ \text{:}\ddot{\text{S}}\text{:} \end{array} + \text{resonance}$
5. H <sub>2</sub> CS 12 $\begin{array}{c} \text{H} \\   \\ \text{C} \\   \\ \text{H} \end{array} = \text{S}\text{:}$	6. CHCl <sub>3</sub> 26 $\begin{array}{c} \text{:}\ddot{\text{Cl}}\text{:} \\   \\ \text{H}-\text{C}-\text{Cl}\text{:} \\   \\ \text{:}\ddot{\text{Cl}}\text{:} \end{array}$
7. SF <sub>2</sub> 20 $\text{:}\ddot{\text{F}}-\ddot{\text{S}}-\ddot{\text{F}}\text{:}$	8. SHF 14 $\text{H}-\ddot{\text{S}}-\ddot{\text{F}}\text{:}$
9. N <sub>2</sub> F <sub>4</sub> 38 $\begin{array}{c} \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \\   \quad   \\ \text{N} \quad \text{N} \\   \quad   \\ \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \end{array}$	10. XeO <sub>4</sub> 32 $\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \\   \quad   \\ \text{Xe} \\   \quad   \\ \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \end{array}$
11. N <sub>2</sub> O <sub>3</sub> (dinitrogen trioxide)* 28e $\text{:}\ddot{\text{O}}=\ddot{\text{N}}-\text{N}=\ddot{\text{O}}\text{:} + \text{resonance}$	12. NO <sub>2</sub> (nitrogen dioxide)* 17e $\text{:}\ddot{\text{O}}=\ddot{\text{N}}-\ddot{\text{O}}\text{:} + \text{resonance}$
13. N <sub>2</sub> O <sub>4</sub> (dinitrogen tetraoxide)* 34 $\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \\   \quad   \\ \text{N} \quad \text{N} \\   \quad   \\ \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \end{array} + \text{resonance}$	14. N <sub>2</sub> O <sub>5</sub> (dinitrogen pentaoxide)* 40e $\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \\   \quad   \\ \text{N} \quad \text{N} \\   \quad   \\ \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \end{array} + \text{resonance}$

\*Problems 11-14 are more challenging and may have more than one correct structure. Please note that if a molecule has an odd number of valence electrons, at least one atom in the molecule will NOT fulfill the octet rule.

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## **Table of Contents Lewis Dot Structure Work Sheet Answer Key**

1. Understanding the eBook Lewis Dot Structure Work Sheet Answer Key
  - The Rise of Digital Reading Lewis Dot Structure Work Sheet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Lewis Dot Structure Work Sheet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Lewis Dot Structure Work Sheet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Lewis Dot Structure Work Sheet Answer Key

- Personalized Recommendations
  - Lewis Dot Structure Work Sheet Answer Key User Reviews and Ratings
  - Lewis Dot Structure Work Sheet Answer Key and Bestseller Lists
5. Accessing Lewis Dot Structure Work Sheet Answer Key Free and Paid eBooks
- Lewis Dot Structure Work Sheet Answer Key Public Domain eBooks
  - Lewis Dot Structure Work Sheet Answer Key eBook Subscription Services
  - Lewis Dot Structure Work Sheet Answer Key Budget-Friendly Options
6. Navigating Lewis Dot Structure Work Sheet Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Lewis Dot Structure Work Sheet Answer Key Compatibility with Devices
  - Lewis Dot Structure Work Sheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Lewis Dot Structure Work Sheet Answer Key
  - Highlighting and Note-Taking Lewis Dot Structure Work Sheet Answer Key
  - Interactive Elements Lewis Dot Structure Work Sheet Answer Key
8. Staying Engaged with Lewis Dot Structure Work Sheet Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Lewis Dot Structure Work Sheet Answer Key
9. Balancing eBooks and Physical Books Lewis Dot Structure Work Sheet Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Lewis Dot Structure Work Sheet Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Lewis Dot Structure Work Sheet Answer Key
- Setting Reading Goals Lewis Dot Structure Work Sheet Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lewis Dot Structure Work Sheet Answer Key

- Fact-Checking eBook Content of Lewis Dot Structure Work Sheet Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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