

# IB CHEMISTRY NOTES

200+ handwritten pages

**Bohr model:**

- Electrons orbit in discrete, concentric shells.
- Electrons in the inner shells are closer to the nucleus and experience a greater attractive force.
- Electrons in the outer shells are further from the nucleus and experience a weaker attractive force.
- Electrons in the outer shells have a higher energy level.

**Energy levels:**

Electrons are arranged in shells. The energy level of an electron is determined by its distance from the nucleus. Electrons in the inner shells have a lower energy level than electrons in the outer shells.

**Spectra:**

The emission spectrum of an element is a series of discrete lines of light. Each line represents a transition of an electron from a higher energy level to a lower energy level. The energy difference between the two levels is equal to the energy of the photon emitted.

**Quantum numbers:**

- Principal quantum number (n):** Determines the energy level and the size of the orbital.
- Angular momentum quantum number (l):** Determines the shape of the orbital.
- Magnetic quantum number (m<sub>l</sub>):** Determines the orientation of the orbital.
- Spin quantum number (m<sub>s</sub>):** Determines the spin of the electron.

**Atomic structure:**

The atom consists of a central nucleus containing protons and neutrons, surrounded by a cloud of electrons. The protons and neutrons are held together by the strong nuclear force, while the electrons are held in the atom by the electrostatic attraction of the protons.

**Energy level diagram:**

The diagram shows the energy levels of an atom. The energy increases as the principal quantum number (n) increases. The energy levels are labeled n=1, n=2, n=3, n=4, n=5, n=6, n=7. The energy difference between n=1 and n=2 is the largest, and it decreases as n increases.

**Spectrum diagram:**

The diagram shows the emission spectrum of an element. The spectrum consists of discrete lines of light. The lines are labeled with their corresponding energy levels. The energy difference between the two levels is equal to the energy of the photon emitted.

**Flowchart:**

The flowchart shows the relationship between the different quantum numbers. The principal quantum number (n) determines the energy level and the size of the orbital. The angular momentum quantum number (l) determines the shape of the orbital. The magnetic quantum number (m<sub>l</sub>) determines the orientation of the orbital. The spin quantum number (m<sub>s</sub>) determines the spin of the electron.

**Additional notes:**

- The energy of a photon is given by  $E = hf$ , where  $E$  is the energy,  $h$  is Planck's constant, and  $f$  is the frequency.
- The wavelength of a photon is given by  $\lambda = \frac{c}{f}$ , where  $\lambda$  is the wavelength,  $c$  is the speed of light, and  $f$  is the frequency.
- The energy difference between two energy levels is equal to the energy of the photon emitted.

Includes detailed handwritten notes for both SL and HL

# Ib Chemistry Sl Study

**Markus Reiher, Alexander Wolf**



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provide a concise comprehensive and complete presentation of this theory This second edition expands on some of the latest developments in this fascinating field The text retains its clear and consistent style allowing for a readily accessible overview of the complex topic It is also self contained building on the fundamental equations and providing the mathematical background necessary While some parts of the text have been restructured for the sake of clarity a significant amount of new content has also been added This includes for example an in depth discussion of the Brown Ravenhall disease of spin in current density functional theory and of exact two component methods and its local variants A strength of the first edition of this textbook was its list of almost 1000 references to the original research literature which has made it a valuable reference also for experts in the field In the second edition more than 100 additional key references have been added most of them considering the recent developments in the field Thus the book is a must have for everyone entering the field as well as for experienced researchers searching for a consistent review

**Updraft Downdraft** Marilyn Crawford,Eleanor

Dougherty,2003-06-11 Ever wonder why there is such a big gap in the number of students who succeed in comparison to those who do not The duality of educational experiences that permeate our secondary schools leads to one group of students heading into post secondary education while others either receive terminal degrees or drop out This updraft downdraft phenomena occurs because time talent and monies tend to drift into the updraft group s educational arena leaving the downdraft students and teachers with less time to teach fewer qualified teachers and fewer supporters This book provides a set of tools to assist schools and communities in reallocating their resources more effectively Using school artifacts educators can collect the data and information needed to make better decisions and leverage what they have equitably The process of collecting analyzing and acting on the data described promotes more strategic decision making about school resources Samples of data and analyses around master schedules student schedules curriculum guides and other artifacts inside the school and district are provided Recommended for anyone seeking ways to understand the problems inside the public schools and ways to find solutions

Molecular Modelling and Synthesis of Nanomaterials Ihsan Boustani,2020-07-14

This book presents nanomaterials as predicted by computational modelling and numerical simulation tools and confirmed by modern experimental techniques It begins by summarizing basic theoretical methods then giving both a theoretical and experimental treatment of how alkali metal clusters develop into nanostructures as influenced by the cluster s magic number of atoms The book continues with a discussion of atomic clusters and nanostructures focusing primarily on boron and carbon exploring in detail the one two and three dimensional structures of boron and carbon and describing their myriad potential applications in nanotechnology from nanocoating and nanosensing to nanobatteries with high borophene capacity The broad discussion of computational modelling as well as the specific applications to boron and carbon make this book an essential reference resource for materials scientists in this field of research

**International Baccalaureate Standard Level Chemistry**

Cameron Lumsden,2015 A Comprehensive Treatise on Inorganic and Theoretical Chemistry: B, Al, Ga, In, Tl, Sc, Ce, and

Rare Earth Metals, C (Part I) Joseph William Mellor, 1924      *Chemical Age*, 1924      **Supplement to Mellor's**  
**Comprehensive Treatise on Inorganic and Theoretical Chemistry** Joseph William Mellor, 1922      Supplement to  
Mellor's Comprehensive Treatise on Inorganic and Theoretical Chemistry: suppl. 3. K, Rb, Cs, Fr Joseph William Mellor, 1922  
Johns Hopkins University Circulars Johns Hopkins University, 1895      **Circulars** Johns Hopkins University, 1889      The  
Johns Hopkins University circular, 1895      *The Johns Hopkins University Circular* Johns Hopkins University, 1886  
*Relativistic Theory of Atoms and Molecules II* Pekka Pyykkö, 1993-11-03 Relativistic effects are of major importance for  
understanding the properties of heavier atoms and molecules This book is still the only comprehensive bibliography on  
related calculations The material is organized by subject into tables containing a concise characterization Together with  
Volume I Lecture Notes in Chemistry Vol 41 ISBN 3 540 17167 3 the literature until 1992 is now covered and 6577  
references with titles are given in the two books The book will provide a convenient reference for theoretical chemists and  
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*Theoretical Chemistry: Ra and Ac families, Be, Mg, Zn, Cd, Hg* Joseph William Mellor, 1923      **Bibliographia**  
**Hopkinsiensis 1876-1893** Johns Hopkins University, 1893      *The Chemical Age*, 1925      **Bulletin of Carson-Newman**  
**College** Carson-Newman College, 1928

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