



illuminati - 2019

Advanced Chemistry Assignment-3B | Inorganic Chemistry

SINGLE CHOICE QUESTIONS

- During refining of pig iron by bessemerization process, which of the following impurity is oxidized last.
 (A) Si (B) Mn (C) S (D) C
- $\text{PCl}_5 + \text{H}_2(\text{g}) \xrightarrow{\text{warm}} \text{X} + \text{Y}$
 If molecular weight of 'X' is greater than that of 'Y' then compound 'X' reacts with :
 (A) $\text{H}_2\text{O}(\text{l})$ (B) Cl_2 gas (C) O_2 gas (D) All
- Incorrect statement for borax is :
 (A) Its aqueous solution acts as buffer
 (B) Its aqueous solution is used to determine strength of dil. HCl using phenolphthalein
 (C) It is used as acidic flux in brazing
 (D) H-bonding is present in borax
- H_2O_2 acts as oxidizing agent with which of the following reagent :
 (A) KMnO_4 in basic medium (B) $\text{Fe}_2(\text{SO}_4)_3$
 (C) I_2 in alkaline medium (D) HClO_4
- Milk of lime + $\text{SOCl}_2 \longrightarrow \text{P} \downarrow + \text{Q}$
 Incorrect statement for compound 'P' and 'Q' is :
 (A) 'Q' produces observable change with litmo
 (B) 'P' dissolves in excess of H_2SO_4
 (C) 'Q' produces precipitate with aqueous solution of KPF
 (D) 'P' reacts with H_2S gas in the presence of dilHCl
- $\text{K}_4[\text{Fe}(\text{CN})_6] + \text{conc. H}_2\text{SO}_4 \xrightarrow{\text{warm}} \text{products}$
 Incorrect statement for products is :
 (A) Gaseous product is used in Mond's process
 (B) Oxidation state of iron remains same
 (C) One of the cations is derived from nonmetals
 (D) Gaseous product obtained is acidic
- $3\text{HgO} \cdot \text{HgCO}_3(\text{s}) + 3\text{PbS}(\text{s}) \xrightarrow{\Delta} \text{Products}$
 Incorrect statements about products is :
 (A) One of the liberated gas can decolourize aqueous KMnO_4 solution.
 (B) Mercury is obtained in elemental state
 (C) Lead is obtained in the form of PbO_2
 (D) PbS is roasted into PbO and SO_2
- Incorrect statement for complex ions $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ and $[\text{MnBr}_4]^{2-}$ is :
 (A) Intensity of colour in $[\text{MnBr}_4]^{2-}$ is less than $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$
 (B) Both have same magnetic moment
 (C) Stability : $[\text{Mn}(\text{H}_2\text{O})_6]^{2+} > [\text{MnBr}_4]^{2-}$
 (D) Valence shell d- orbitals of Mn^{2+} are used in its hybridization in $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$

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