

AP Chemistry Chapter 8 Answers – Zumdahl

8.87

The first step always is to draw a valid Lewis structure when predicting molecular structure. When resonance is possible, only one of the possible resonance structures is necessary to predict the correct structure because all resonance structures give the same structure. The Lewis structures are in Exercises 8.67 and 8.73. The structures and bond angles for each follow.

- 8.67
- HCN: linear, 180°
 - PH_3 : trigonal pyramid, $< 109.5^\circ$
 - CHCl_3 : tetrahedral, 109.5°
 - NH_4^+ : tetrahedral, 109.5°
 - H_2CO : trigonal planar, 120°
 - SeF_2 : V-shaped or bent, $< 109.5^\circ$
 - CO_2 : linear, 180°

h and I_2O_2 and HBr are both linear, but there is no bond angle in either.

Note: PH_3 and SeF_2 both have lone pairs of electrons on the central atom which result in bond angles that are something less than predicted from a tetrahedral arrangement (109.5°). However, we cannot predict the exact number. For the solutions manual, we will insert a less than sign to indicate this phenomenon. For bond angles equal to 120° , the lone pair phenomenon isn't as significant as compared to smaller bond angles. For these molecules, e.g., NO_2^- , we will insert an approximate sign in front of the 120° to note that there may be a slight distortion from the VSEPR bond angle.

- 8.73
- NO_2^- : V-shaped, approximately 120° ; NO_3^- : trigonal planar, 120°
 N_2O_4 : trigonal planar, 120° about both N atoms
 - OCN^- , SCN^- , and N_3^- are all linear with 180° bond angles.

8.89

From the Lewis structures (see Exercise 8.71), Br_3^- would have a linear molecular structure, ClF_3 would have a T-shaped molecular structure and SF_4 would have a see-saw molecular structure.

For example, consider ClF_3 (28 valence electrons): The central Cl atom is surrounded by 5 electron pairs, which requires a trigonal bipyramidal geometry. Since there are 3 bonded atoms and 2 lone pairs of electrons about Cl, we describe the molecular structure of ClF_3 as T-shaped with predicted bond angles of about 90° . The actual bond angles would be slightly less than 90° due to the strong repulsive effect of the lone pair electrons as compared to the bonding electrons.

8.91

- SeO_3 has a trigonal planar molecular structure with all bond angles equal to 120° . Note that any one of the resonance structures could be used to predict molecular structure and bond angles.

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L Cohen



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