

## Covalent Bonding Review Sheet – Dec 12, 2012

Name \_\_\_\_\_ class period \_\_\_\_\_

- Given the formula, a compound can be named using conventional systems that include Greek prefixes and Roman numeral where appropriate.
- Differences in electronegativity values can be used to predict where a bond fits on the continuum between ionic and covalent bonds.
- Since most compounds contain multiple bonds, a substance may contain more than one type of bond.

### Section 1: Naming and Formula Writing

Write the formulas for the following covalent compounds

1. Nitrogen Monoxide NO
2. Dinitrogen Monoxide N<sub>2</sub>O
3. Nitrogen Dioxide NO<sub>2</sub>
4. Dinitrogen Trioxide N<sub>2</sub>O<sub>3</sub>
5. Dinitrogen Tetraoxide N<sub>2</sub>O<sub>4</sub>
6. Dinitrogen Pentaoxide N<sub>2</sub>O<sub>5</sub>
7. carbon dioxide CO<sub>2</sub>
8. phosphorus triiodide PI<sub>3</sub>
9. sulfur dichloride SCl<sub>2</sub>
10. nitrogen trifluoride NF<sub>3</sub>

Write names from the following formulas

11. N<sub>2</sub>F<sub>4</sub> dinitrogen tetrafluoride
12. SCl<sub>4</sub> sulfur tetrachloride
13. ClF<sub>3</sub> chlorine trifluoride
14. SiO<sub>2</sub> silicon dioxide
15. P<sub>4</sub>O<sub>10</sub> tetraphosphorus decaoxide
16. SiF<sub>4</sub> silicon tetrafluoride
17. H<sub>2</sub>O dihydrogen monoxide
18. Br<sub>2</sub> dibromide or bromine
19. HBr hydrogen monobromide
20. N<sub>2</sub>S<sub>3</sub> dinitrogen trisulfide

# Concept Review Section Covalent Bonds Answer Key

**Katrin Zwirgmaier**



**Concept Review Section Covalent Bonds Answer Key:**

## Unveiling the Energy of Verbal Art: An Psychological Sojourn through **Concept Review Section Covalent Bonds Answer Key**

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### **Table of Contents Concept Review Section Covalent Bonds Answer Key**

1. Understanding the eBook Concept Review Section Covalent Bonds Answer Key
  - The Rise of Digital Reading Concept Review Section Covalent Bonds Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Section Covalent Bonds Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Concept Review Section Covalent Bonds Answer Key
  - User-Friendly Interface

4. Exploring eBook Recommendations from Concept Review Section Covalent Bonds Answer Key
  - Personalized Recommendations
  - Concept Review Section Covalent Bonds Answer Key User Reviews and Ratings
  - Concept Review Section Covalent Bonds Answer Key and Bestseller Lists
5. Accessing Concept Review Section Covalent Bonds Answer Key Free and Paid eBooks
  - Concept Review Section Covalent Bonds Answer Key Public Domain eBooks
  - Concept Review Section Covalent Bonds Answer Key eBook Subscription Services
  - Concept Review Section Covalent Bonds Answer Key Budget-Friendly Options
6. Navigating Concept Review Section Covalent Bonds Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Concept Review Section Covalent Bonds Answer Key Compatibility with Devices
  - Concept Review Section Covalent Bonds Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Concept Review Section Covalent Bonds Answer Key
  - Highlighting and Note-Taking Concept Review Section Covalent Bonds Answer Key
  - Interactive Elements Concept Review Section Covalent Bonds Answer Key
8. Staying Engaged with Concept Review Section Covalent Bonds Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Concept Review Section Covalent Bonds Answer Key
9. Balancing eBooks and Physical Books Concept Review Section Covalent Bonds Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Concept Review Section Covalent Bonds Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Concept Review Section Covalent Bonds Answer Key
  - Setting Reading Goals Concept Review Section Covalent Bonds Answer Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Concept Review Section Covalent Bonds Answer Key
  - Fact-Checking eBook Content of Concept Review Section Covalent Bonds Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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