

**I- 1)** X un élément qui appartient au groupe IIA et à la deuxième période, sa couche de valence sera de type  $ns^2$  avec  $n = 2$ , donc  $2s^2$ .

Structure électronique de X :  $1s^2 2s^2$ . Son numéro atomique  $Z = 4$ .

**2)**  $X^{2+}$  est l'ion hydrogénoïde de X.

**3)** Constante de Rhydberg relative à  $X^{2+}$  est :  $K = Z^2 \cdot R_{H1} = 4^2 \cdot 1,09678 \cdot 10^7 = 17,54848 \cdot 10^7 \text{ m}^{-1}$ .

**4)** Longueur d'onde de la radiation capable d'ioniser  $X^{2+}$  est :

$$\frac{1}{\lambda_{\text{ion}}} = K \left( \frac{1}{1^2} - \frac{1}{\infty^2} \right) \Rightarrow \lambda_{\text{ion}} = \frac{1}{K} = \frac{1}{17,54848 \cdot 10^7} = 0,05698 \cdot 10^{-7} \text{ m} = 56,98 \text{ \AA}$$

$$5) \Delta E = \frac{hc}{\lambda} = \frac{6,626 \cdot 10^{-34} \cdot 3 \cdot 10^8}{56,9 \cdot 10^{-10}} = 0,348 \cdot 10^{-16} = 348 \cdot 10^{-19} \text{ J} = 2,16 \text{ eV}$$

**II- 1)**  $n$  : nombre quantique principal, il détermine le niveau d'énergie ou la couche associée à l'électron.  $n$  ne peut prendre que des valeurs entières :  $n(\text{entier}) \geq 1$ .

$l$  : nombre quantique secondaire ou azimutal, il détermine la forme générale de l'orbitale c'est-à-dire la région dans laquelle l'électron se déplace (ou configuration spatiale) :  $0 \leq l(\text{entier}) \leq n-1$

$m$  : nombre quantique magnétique,  $m$  est un nombre d'entier qui détermine le nombre d'orbitales associées à un niveau d'énergie donné.

Exemple : pour  $n=3$ ,  $l$  peut être 0, 1 ou 2. Si  $l=0$ ,  $m$  peut être 0. Si  $l=1$ ,  $m$  peut être -1, 0 ou 1. Si  $l=2$ ,  $m$  peut être -2, -1, 0, 1 ou 2.

Exemple : pour  $n=2$ ,  $l$  peut être 0 ou 1. Si  $l=0$ ,  $m$  peut être 0. Si  $l=1$ ,  $m$  peut être -1, 0 ou 1.

Exemple : pour  $n=1$ ,  $l$  peut être 0. Si  $l=0$ ,  $m$  peut être 0.

Exemple : pour  $n=4$ ,  $l$  peut être 0, 1, 2 ou 3. Si  $l=0$ ,  $m$  peut être 0. Si  $l=1$ ,  $m$  peut être -1, 0 ou 1. Si  $l=2$ ,  $m$  peut être -2, -1, 0, 1 ou 2. Si  $l=3$ ,  $m$  peut être -3, -2, -1, 0, 1, 2 ou 3.

Exemple : pour  $n=5$ ,  $l$  peut être 0, 1, 2, 3 ou 4. Si  $l=0$ ,  $m$  peut être 0. Si  $l=1$ ,  $m$  peut être -1, 0 ou 1. Si  $l=2$ ,  $m$  peut être -2, -1, 0, 1 ou 2. Si  $l=3$ ,  $m$  peut être -3, -2, -1, 0, 1, 2 ou 3. Si  $l=4$ ,  $m$  peut être -4, -3, -2, -1, 0, 1, 2, 3 ou 4.

Exemple : pour  $n=6$ ,  $l$  peut être 0, 1, 2, 3, 4 ou 5. Si  $l=0$ ,  $m$  peut être 0. Si  $l=1$ ,  $m$  peut être -1, 0 ou 1. Si  $l=2$ ,  $m$  peut être -2, -1, 0, 1 ou 2. Si  $l=3$ ,  $m$  peut être -3, -2, -1, 0, 1, 2 ou 3. Si  $l=4$ ,  $m$  peut être -4, -3, -2, -1, 0, 1, 2, 3 ou 4. Si  $l=5$ ,  $m$  peut être -5, -4, -3, -2, -1, 0, 1, 2, 3, 4 ou 5.

Exemple : pour  $n=7$ ,  $l$  peut être 0, 1, 2, 3, 4, 5 ou 6. Si  $l=0$ ,  $m$  peut être 0. Si  $l=1$ ,  $m$  peut être -1, 0 ou 1. Si  $l=2$ ,  $m$  peut être -2, -1, 0, 1 ou 2. Si  $l=3$ ,  $m$  peut être -3, -2, -1, 0, 1, 2 ou 3. Si  $l=4$ ,  $m$  peut être -4, -3, -2, -1, 0, 1, 2, 3 ou 4. Si  $l=5$ ,  $m$  peut être -5, -4, -3, -2, -1, 0, 1, 2, 3, 4 ou 5. Si  $l=6$ ,  $m$  peut être -6, -5, -4, -3, -2, -1, 0, 1, 2, 3, 4, 5 ou 6.



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**Annelies Wilder-Smith**



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