

Thermodynamics worksheet

Standard bond enthalpy

- 1) Calculate the bond energy of Cl-Cl bond from the following data



Also the bond enthalpies of C-H, C-Cl, H-Cl are 413, 326, 431 KJ/mol

respectively.

- 2) Propose the structure of $\text{C}_2\text{H}_2\text{Cl}_2$. Calculate the change in enthalpy of the reaction.



Given that average bond enthalpies are C-C, C-H, C-Cl, C=C and C=C are

347, 413, 326, 614, 614 KJ/mol respectively.

- 3) Write the reaction $\text{C}_2\text{H}_2 + \text{Cl}_2 \rightarrow \text{C}_2\text{H}_2\text{Cl}_2 + \text{C}_2\text{H}_4$

calculate the bond energy of C-Cl. Given bond energies of C-H, C-H, C-H and C-H are 413, 413, 413 and 614 KJ/mol respectively.

- 4) Calculate the bond enthalpy of H-Cl given that the bond enthalpies of H₂ and Cl₂ are 436 KJ/mol and 242 KJ/mol and enthalpy of formation for HCl is

-92 KJ/mol.

- 5) Calculate the bond energy of C-C bond given that the heat of formation of methane, heat of combustion of carbon, heat of dissociation of H₂ are

-75 KJ, 393 KJ, 436 KJ/mol respectively.



- 6) Calculate the enthalpy change for the reaction $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$.

Given the bond enthalpies of H-H, Cl-Cl, H-Cl is 436, 242 and 431 KJ/mol respectively.

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