
INSTRUMENT ENGINEERS' HANDBOOK

Fourth Edition

Process Measurement and Analysis

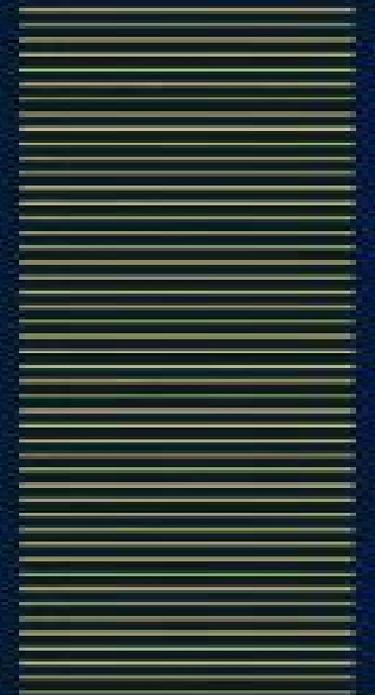
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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation: $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$. 18g H_2O is your theoretical yield so percent yield = $\frac{14.9 \text{ g } \text{H}_2\text{O}}{18 \text{ g } \text{H}_2\text{O}} \times 100 = 82.7\%$. Problems 6 & 8 are similar to problem 5.

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produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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