

Physical Science Chapter 7 Study Guide

- **Molecules**—groups of two or more atoms linked by chemical bonds to form distinct units
- **Molecular mass**—the sum of the mass of the atoms in a molecule
- **Isomers**—molecules with the same number and types of atoms but in a different arrangement
- **Elements**—substances composed of a single type of atom
- **Compounds**—substances composed of different types of atoms bonded together
- **Mixtures**—different kinds of atoms that are physically mixed but not chemically united
 - **Solution**—a mixture in which the ingredients are uniformly mixed
 - **Solute**—that which is dissolved in a substance
 - **Solvent**—the substance into which something is dissolved
 - **Solubility**—the ability of one substance to dissolve in another substance
 - **Precipitate**—solid particles that come out of a solution
 - **Heterogeneous mixture**—substances which are mixed but not completely or uniformly mixed
 - **Colloid**—a mixture that contains tiny suspended clumps of particles
- **Chemical bonds**—the attraction between atoms making up a molecule
 - **Covalent**—a bond which results from a sharing of electrons
 - **Single covalent bonds** means sharing one pair of electrons; double means two pairs, etc.
 - **Lewis structure**—a diagram that shows all of the atoms in a molecule and the valence electrons around each atom
 - **Polar vs. nonpolar covalent bonds**
 - **Polar bond**—the one atom has a stronger pull on the electrons (this is called its electronegativity) and therefore “hogs” the electrons most of the time, creating a negative electrical force around it and leaving the other atom with a positive electrical force
 - **Nonpolar bond**—the atoms in the molecule have a basically equal pull on the electrons; therefore the charge remains balanced
 - **Ionic bond**—attraction between oppositely charged atoms; occurs when one atom transfers electrons to another atom
 - **Formula unit**—The simplest ratio of atoms in a nonmolecular compound
 - **Metallic bond**—occurs between metal elements and involves the sharing of electrons around many nuclei (think Hutterite colony)
- **Chemical formulas**
 - **Molecular formula**—shows the exact number and kind of atoms in a molecule
 - **Structural formula**—shows the general arrangement of atoms in a molecule
 - **Empirical formula**—shows only the simplest ratio of atoms in molecule

Physical Science Chapter 7 Study Guide Answers

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Resources for Teaching

Middle School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science centers museums and zoos where teachers can take middle school students for interactive science experiences Another chapter describes nearly 140 professional associations and U S government agencies that offer resources and assistance Authoritative extensive and thoroughly indexed and the only guide of its kind Resources for Teaching Middle School Science will be the most used book on the shelf for science teachers school administrators teacher trainers science curriculum specialists advocates of hands on science teaching and concerned parents

Focus on

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 - Chapter 3: Physical Science Chapter 7 Study Guide Answers in Everyday Life
 - Chapter 4: Physical Science Chapter 7 Study Guide Answers in Specific Contexts
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