

ANSWERS

Nuclear Chemistry Review Questions

1. Compare the mass, charge, and penetration power of alpha, beta and gamma radiation.

	Mass	Charge	Penetration
Alpha	4	+	Lowest
Beta	0	-	Middle
Gamma	0	0	Highest

2. In 5.49 seconds, 1.20g of argon-35 decays to leave only 0.15g. What is the half-life?

1.20g \rightarrow 0.60g \rightarrow 0.30g \rightarrow 0.15g , therefore it took 3 half-life periods.

$$5.49 \div 3 = 1.83 \text{ sec}$$

3. How many days does it take 16.0g of gold-198 to decay to 1.0g?

16.0g \rightarrow 8.0g \rightarrow 4.0g \rightarrow 2.0g \rightarrow 1.0g , therefore 4 half-life periods

From Ref Table N, 1 half-life = 2.69d

$$2.69\text{d} \times 4 = 10.76 \text{ days}$$

4. Why does an atom undergo radioactive decay?

Because it has an unstable nucleus

5. Compare and contrast fission and fusion. Why is fission used for electrical and fusion is not?

Fission=Division, the splitting of large unstable atoms(nucleus)

Fusion=joining together, small isotopes of H, creates way more energy

6. How is nuclear chemistry applied to medicine? What type of radioisotope must be used when it is injected into a person?

Radioisotopes are used to treat and diagnose illnesses

If injected they must have short half-lives and quickly removed from the body

7. Write out the nuclear decay of radium-226.



8. What is different about nuclear reactions when compared to a chemical reaction?

Nuclear reactions transform atoms into other atoms, chemical reactions involved the rearrangement of electrons

9. At what point do all isotopes of an element become radioactive?

Up to and including atomic # 83

Chapter 25 Nuclear Chemistry Worksheet Answers

LL Leslie



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Table of Contents Chapter 25 Nuclear Chemistry Worksheet Answers

1. Understanding the eBook Chapter 25 Nuclear Chemistry Worksheet Answers
 - The Rise of Digital Reading Chapter 25 Nuclear Chemistry Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 25 Nuclear Chemistry Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 25 Nuclear Chemistry Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 25 Nuclear Chemistry Worksheet Answers
 - Personalized Recommendations
 - Chapter 25 Nuclear Chemistry Worksheet Answers User Reviews and Ratings
 - Chapter 25 Nuclear Chemistry Worksheet Answers and Bestseller Lists
5. Accessing Chapter 25 Nuclear Chemistry Worksheet Answers Free and Paid eBooks
 - Chapter 25 Nuclear Chemistry Worksheet Answers Public Domain eBooks
 - Chapter 25 Nuclear Chemistry Worksheet Answers eBook Subscription Services
 - Chapter 25 Nuclear Chemistry Worksheet Answers Budget-Friendly Options
6. Navigating Chapter 25 Nuclear Chemistry Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 25 Nuclear Chemistry Worksheet Answers Compatibility with Devices
 - Chapter 25 Nuclear Chemistry Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 25 Nuclear Chemistry Worksheet Answers
 - Highlighting and Note-Taking Chapter 25 Nuclear Chemistry Worksheet Answers
 - Interactive Elements Chapter 25 Nuclear Chemistry Worksheet Answers
8. Staying Engaged with Chapter 25 Nuclear Chemistry Worksheet Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 25 Nuclear Chemistry Worksheet Answers
9. Balancing eBooks and Physical Books Chapter 25 Nuclear Chemistry Worksheet Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 25 Nuclear Chemistry Worksheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 25 Nuclear Chemistry Worksheet Answers
- Setting Reading Goals Chapter 25 Nuclear Chemistry Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 25 Nuclear Chemistry Worksheet Answers
- Fact-Checking eBook Content of Chapter 25 Nuclear Chemistry Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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