

## 11. Chapter 11: Thermodynamics

1. An adiabatic expansion process for an ideal gas is

- |                                   |   |
|-----------------------------------|---|
| (A) $TV = \text{constant}$        | (D) $TV^\gamma = \text{constant}$ (where $\gamma > 1$ ) |
| (B) $PV = \text{constant}$        | (E) $PV^\gamma = \text{constant}$ (where $\gamma > 1$ ) |
| (C) $TV^\gamma = \text{constant}$ |   |

2. The adiabatic index  $\gamma$  for a diatomic gas is the ratio of its

- |  |  |
|--|--|
| (A) constant volume molar specific heat to its constant pressure molar specific heat | (D) constant volume molar specific heat to its constant pressure molar specific heat |
| (B) constant pressure molar specific heat to its constant volume molar specific heat | (E) constant volume molar specific heat to its constant pressure molar specific heat |
| (C) constant pressure molar specific heat to its constant volume molar specific heat |  |

3. An adiabatic process for an ideal gas is described by  $PV^\gamma = \text{constant}$ . Calculate the molar specific heat  $c_{\text{eff}}$  for the expansion of an ideal diatomic gas from 300 K to 600 K.

- (A) 1.5R (B) 2.0R (C) 2.5R (D) 3.0R (E) 3.5R

Answer: (D)

4. A diatomic molecule (1.0 g) is heated at 300 K. In the process, 500 J of heat is added to the gas. The change in the internal energy of the gas is

- (A) 400 J (B) 300 J (C) 200 J (D) 100 J (E) 50 J

Answer: (D)

5. A 1.00 mole of water at 300 K is heated to 350 K. What is the change in the internal energy of the water? The specific heat of water is 4.18 J/g·K.

- (A) 209 J (B) 418 J (C) 836 J (D) 1672 J (E) 3344 J

Answer: (D)

6. A 1.00 g sample of 1.00 mol of  $\text{H}_2\text{O}$  is heated. The molar specific heat of  $\text{H}_2\text{O}$  at constant volume is 12.47 J/mol·K. What is the final temperature of the water? The specific heat of water is 4.18 J/g·K.

- (A) 300 K (B) 350 K (C) 400 K (D) 450 K (E) 500 K

Answer: (E)

7. A piece of copper with a mass of 100 g has a heat capacity of 100 J/K. What is the specific heat of copper?

- |                |                |
|----------------|----------------|
| (A) 1.00 J/g·K | (D) 1.00 J/g·K |
| (B) 1.00 J/g·K | (E) 1.00 J/g·K |
| (C) 1.00 J/g·K |                |

Answer: (E)

# Chapter 6 Thermochemistry Test

**Jessica J Manson**



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