

## CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron III sulfate, the roman numeral tells us **C**
  - How many atoms of Fe are in one formula unit
  - How many sulfate ions can be attached to the iron atom
  - The charge on each Fe ion
  - The total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to **c**
  - Change the number of moles represented in the formula
  - Change the charges on the other ions in the compound
  - Change the formula so that it no longer represents the compound
  - Have no effect on the formula
- Derive a generalization for when an acid name will end in the suffix **-ic** or **-ous**  
**If ending is -ic, the anion ended in -ate; if ending is -ous, the anion ended in -ite**
- Derive a generalization for when an acid name will begin with the prefix **hydro-**.  
**If prefix is hydro-, the anion ended in -ide. This is most likely a monatomic ion.**
- Fill in the blanks in the table below

| Compound Name               | Formula        |
|-----------------------------|----------------|
| Aluminum sulfide            | $Al_2S_3$      |
| Aluminum sulfate            | $Al_2(SO_4)_3$ |
| Lead II chloride            | $PbCl_2$       |
| Ammonium phosphite          | $(NH_4)_2PO_3$ |
| Hydroiodic acid             | $HCl$          |
| Dinitrogen pentoxide        | $N_2O_5$       |
| Copper II carbonate         | $CuCO_3$       |
| Sodium sulfide              | $Na_2S$        |
| Ammonium phosphate          | $(NH_4)_3PO_4$ |
| Tin IV sulfite              | $Sn(SO_3)_2$   |
| Nitrous acid                | $HNO_2$        |
| Magnesium perchlorate       | $Mg(ClO_4)_2$  |
| Iron II nitrate             | $Fe(NO_3)_2$   |
| Iron III nitrate            | $Fe(NO_3)_3$   |
| Cobalt II oxide             | $CoO$          |
| Sulfur trioxide             | $SO_3$         |
| Boron pentafluoride         | $BF_5$         |
| Tetraphosphorus octasulfide | $P_4S_8$       |
| Tetraiodide nonoxide        | $I_4O_9$       |

- Name the following compound and identify as molecular, ionic or acid
  - $Ca(HCO_3)_2$  **calcium bicarbonate - I**
  - $Fe(ClO_2)_3$  **Iron III chlorite - I**
  - $Fe(ClO_3)_2$  **Iron II chlorate - I**
  - $N_2O_5$  **dinitrogen pentoxide - M**

# Chapter 9 Chemical Names Formulas Answers Key

**Michael Seilmaier**



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