

Stereochemistry

You might think that compounds of the same molecular formulas are the same, but that is always not true. There are two reasons for that: the first is that atoms can be arranged differently, the second is that atoms can orient differently in space so that the molecules acquire different 3D structures [configurations]. Look at these two examples:



Figure 1



Figure 2

In the first examples both compounds have the formula $\text{C}_2\text{H}_6\text{O}$, but the compound on the left is an ether while the right one is an alcohol and both compounds exhibit different physical and chemical properties.

In the second example atoms have the same connectivity but they differ in their spatial arrangement. The S- configuration is an anti-inflammatory drug while the R-configuration is a liver toxin. The difference is that the methyl group is drawn dashed while that in the R-configuration is drawn as a wedge. Organic chemists usually use dashed lines to show that the

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