

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ **lattice energy** _____ a. ionic bond
 _____ **bond energy** _____ b. covalent bond
 _____ **enthalpy of vaporization** _____ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ **ionic** _____ a. H—F _____ **nonpolar covalent** _____ d. H—H
 _____ **ionic** _____ b. Na—Cl _____ **polar covalent** _____ e. H—C
 _____ **polar covalent** _____ c. H—O _____ **polar covalent** _____ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H_2S and H_2O have similar structures and their central atoms belong to the same group. Yet H_2S is a gas at room temperature and H_2O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H_2O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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RS Peters



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Table of Contents Chapter 8 Covalent Bonding Work Answers Prentice Hall

1. Understanding the eBook Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - The Rise of Digital Reading Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Personalized Recommendations
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall User Reviews and Ratings
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall and Bestseller Lists
5. Accessing Chapter 8 Covalent Bonding Work Answers Prentice Hall Free and Paid eBooks
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Public Domain eBooks
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall eBook Subscription Services
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Budget-Friendly Options

6. Navigating Chapter 8 Covalent Bonding Work Answers Prentice Hall eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Compatibility with Devices
 - Chapter 8 Covalent Bonding Work Answers Prentice Hall Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Highlighting and Note-Taking Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Interactive Elements Chapter 8 Covalent Bonding Work Answers Prentice Hall
8. Staying Engaged with Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Covalent Bonding Work Answers Prentice Hall
9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Work Answers Prentice Hall
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Setting Reading Goals Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Work Answers Prentice Hall
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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