

14. Circle the pair shown below, that could act as a buffer in solution. Explain your choice.

a.  $\text{HCl}$  &  $\text{NaOH}$

b.  $\text{C}_6\text{H}_5\text{O}_2$  &  $\text{C}_6\text{H}_5\text{O}_2^-$

15. Answer the following questions related to buffers:

a. What two components are required in every buffered solution?

b. Provide an example of a buffered solution.

c. Which component will counteract the addition of an acid to the system? A base?

*Skill* – The student will be able to calculate an acid – base stoichiometry problem.

16. Determine the volume of 0.050 M  $\text{HCl}$  required to titrate 25.0 mL of 0.445 M  $\text{Sr}(\text{OH})_2$ .

17. Determine the concentration of sodium hydroxide that is required if 35.0 mL is necessary to neutralize 57.2 mL of 0.425 M perchloric acid during a titration.

## Acids & Bases: Assessment Review

Name \_\_\_\_\_

*Skill* – The student will be able to identify the acid, base, conjugate acid, and conjugate base in an acid – base reaction.

1. Write a chemical equation showing how  $\text{HCO}_3^-$  behaves as an acid when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

# Chapter 8 Solutions Acids Bases Assessment Answer Key

**J Rink**



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