

CHAPTER 8 – ACIDS & BASES

Section 8.1-Understanding Acids & Bases

Properties of Pure & Aqueous Substances

Substance	Conductivity			Liquid
	Solid	Liquid	Aqueous	Aqueous
Most molecular compounds	No	No	No	No Effect
Most Ionic compounds	No	Yes	Yes	No Effect
Acids	No	No	Yes	Blue- → Red
Bases	No	Yes	Yes	Red- → Blue

- $\text{NaOH}_{(s)} \rightarrow \text{Na}^+_{(aq)} + \text{OH}^-_{(aq)}$ similar to ionic compounds
- acids in liquids are still showing acid's molecular nature
- something happens when a solid acid is dissolved in water i.e.
 $\text{HCl}_{(g)} \rightarrow \text{H}^+_{(aq)} + \text{Cl}^-_{(aq)}$
- Arrhenius hypothesized that bases will dissociate into individual positive & negative ions
e.g. $\text{Ba}(\text{OH})_{2(s)} \rightarrow \text{Ba}^{+2}_{(aq)} + 2\text{OH}^{-1}_{(aq)}$
- Arrhenius also theorized that acids first dissolve as individual molecules and then ionize into H^+ and X^- in solution
e.g. $\text{HCl}_{(g)} \rightarrow \text{H}^+_{(aq)} + \text{Cl}^-_{(aq)}$

Sample Problem 1

Chapter 8 Solutions Acids Bases Section 82 Answers

Michael Rajnik



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