

## Chapter Test A

### Chapter: Chemical Equations and Reactions

Do the work provided below the letter of the item or phrase that best completes each statement or best answers each question.

1. The mass number of an atom is the sum of the number of \_\_\_\_\_ and the number of \_\_\_\_\_ in the nucleus of the atom.
  - a. electrons; neutrons
  - b. protons; neutrons
  - c. protons; electrons
  - d. neutrons; electrons
2. An element is described by a chemical equation:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ .
  - a. oxidized hydrogen is a product.
  - b. oxidized two particles.
  - c. reduced is a reactant.
  - d. both a and c.
3. Which of the following relationships is always true for a balanced chemical equation?
  - a. The number of atoms of each element is the same on both sides of the equation.
  - b. All of the coefficients are the same.
  - c. The number of molecules is the same on both sides of the equation.
  - d. The number of molecules is the same on both sides of the equation.
4. There is a negative enthalpy change for a chemical reaction. Which of the following is true?
  - a. The reactants are at a higher energy level than the products.
  - b. The products are at a higher energy level than the reactants.
  - c. The reactants are at a higher energy level than the products.
  - d. All of the above.
5. The number of chemical equations that can be balanced is \_\_\_\_\_ .
  - a. equal the number.
  - b. equal the number of products.
  - c. equal the number of reactants.
  - d. equal the coefficients.

# Chapter 9 Modern Chemistry Test

**Marcel A. Müller**



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