

Section 12.1

OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1. $\text{Fe}^{2+} + \text{MnO}_4^-$
 - (a) in acidic solution
 - (b) in basic solution
2. $\text{Fe}^{2+} + \text{H}_2\text{O}_2$
 - (a) in acidic solution
 - (b) in basic solution
3. $\text{Fe}^{2+} + \text{H}_2\text{O}_2$
 - (a) in acidic solution
 - (b) in basic solution
4. $\text{Fe}^{2+} + \text{H}_2\text{O}_2$
 - (a) in acidic solution
 - (b) in basic solution
5. $\text{Fe}^{2+} + \text{H}_2\text{O}_2$
 - (a) in acidic solution
 - (b) in basic solution

6. Write the net ionic equation for the following redox reaction:



a. Balance the equation:



b. What is the oxidation of Mn, or O?

c. How many electrons are transferred in the reaction?

d. Write a balanced equation for the reaction in acidic solution.

e. Write a balanced equation for the reaction in basic solution.

Chapter 9 Review Stoichiometry Answers Section 1

Thomas A. Evangelist



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