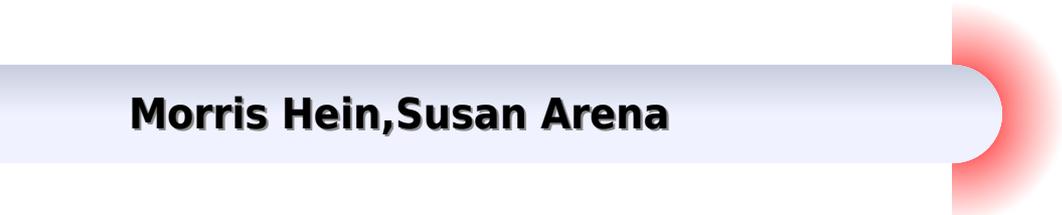


Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
 - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
 - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
 - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} = \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} = \frac{1 \text{ mol CaO}}{1 \text{ mol Ca}}$ $\frac{2 \text{ mol Ca}}{2 \text{ mol CaO}} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol CaO}} = \frac{2 \text{ mol CaO}}{2 \text{ mol CaO}} \cdot \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2}$ *Remember the ratios can be reduced as well!!!*
 - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$ $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} = \frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$ $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$ $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} = \frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$ $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$ $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} = \frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$
- $\frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{105.99 \text{ g Na}_2\text{CO}_3} = \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \cdot \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} = \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} = \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3}$
- Molar Mass
 - Definition: the mass of one mole of a substance
 - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen \rightarrow Water
 - 5.0 mol H₂
 - 2.5 mol O₂
- Ethane combustion reaction
 - 15.8 mol O₂
 - 9.00 mol CO₂ & 13.5 mol H₂O
- 575 g Na & 886 g Cl₂
- Iron production reaction
 - 75.1 mol CO
 - 50.1 mol Fe & 75.1 mol CO₂
- 87,390 g CO or 8.739×10^4 g CO & 12,610 g H₂ or 1.261×10^4 g H₂
- Lightning reaction of N₂ & O₂
 - 1,100 g NO₂ or 1.10×10^3 g NO₂
 720. g NO or 7.20×10^2 g NO
- CO₂ removal in spacecraft
 - 42.04 mol NaOH
 - 21.02 mol Na₂CO₃ & 21.02 mol H₂O
- Double replacement AgNO₃ & NaBr
 - 463 g NaBr
 - 845 g AgBr
- Soda-acid fire extinguisher
 - 3.058 mol NaHCO₃
 - 3.058 mol CO₂, 1.529 mol Na₂SO₄, 3.058 mol H₂O
- Sulfuric acid-sodium hydroxide neutralization reaction
 - H₂SO₄ + 2 NaOH \rightarrow Na₂SO₄ + 2 H₂O
 - 37 g H₂SO₄
 - 53 g Na₂SO₄ & 14 g H₂O
- Copper & silver nitrate reaction
 - 0.0104 mol Cu(NO₃)₂
 - 0.0104 mol Cu & 0.0209 mol AgNO₃
- Aspirin production
 - 13.5 kg C₉H₈O₄
 - 7.66 kg C₉H₈O₃
 - 4.29 L HC₂H₃O₂
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

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Table of Contents Chapter 9 Review Stoichiometry Section 2 Answers

1. Understanding the eBook Chapter 9 Review Stoichiometry Section 2 Answers
 - The Rise of Digital Reading Chapter 9 Review Stoichiometry Section 2 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Review Stoichiometry Section 2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Review Stoichiometry Section 2 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Review Stoichiometry Section 2 Answers
 - Personalized Recommendations
 - Chapter 9 Review Stoichiometry Section 2 Answers User Reviews and Ratings
 - Chapter 9 Review Stoichiometry Section 2 Answers and Bestseller Lists
5. Accessing Chapter 9 Review Stoichiometry Section 2 Answers Free and Paid eBooks

- Chapter 9 Review Stoichiometry Section 2 Answers Public Domain eBooks
 - Chapter 9 Review Stoichiometry Section 2 Answers eBook Subscription Services
 - Chapter 9 Review Stoichiometry Section 2 Answers Budget-Friendly Options
6. Navigating Chapter 9 Review Stoichiometry Section 2 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Review Stoichiometry Section 2 Answers Compatibility with Devices
 - Chapter 9 Review Stoichiometry Section 2 Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Review Stoichiometry Section 2 Answers
 - Highlighting and Note-Taking Chapter 9 Review Stoichiometry Section 2 Answers
 - Interactive Elements Chapter 9 Review Stoichiometry Section 2 Answers
 8. Staying Engaged with Chapter 9 Review Stoichiometry Section 2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Review Stoichiometry Section 2 Answers
 9. Balancing eBooks and Physical Books Chapter 9 Review Stoichiometry Section 2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Review Stoichiometry Section 2 Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 9 Review Stoichiometry Section 2 Answers
 - Setting Reading Goals Chapter 9 Review Stoichiometry Section 2 Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 9 Review Stoichiometry Section 2 Answers
 - Fact-Checking eBook Content of Chapter 9 Review Stoichiometry Section 2 Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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