

EXPERIMENT A5: TYPES OF REACTIONS

Learning Outcomes

Upon completion of this lab, the student will be able to:

- 1) Examine different types of chemical reactions.
- 2) Express chemical equations in molecular, ionic, and net-ionic forms.
- 3) Predict solubility rules for some ionic substances in water.

Introduction

There are several different types of chemical reactions such as: combustion, precipitation, acid-base, decomposition, oxidation-reduction etc. Some evidences for the occurrence of a chemical reaction include one or more of the following:

- Formation of a precipitate
- Color change
- Gas formation
- Temperature change

Three different types of chemical reactions will be considered in this experiment: 1) precipitation 2) acid-base and 3) gas formation. For each type of reaction, it is important to learn to write the molecular equation, total ionic equation, and net ionic equation. The observations from these reactions will enable the development of the solubility rules for simple ionic substances.

Precipitation Reactions

When aqueous solutions of two inorganic salts are combined, a type of reaction referred to as *metathesis reaction* takes place. These reactions are also referred to as double replacement or double displacement reactions. In this reaction the cation of the first solution combines with the anion of the second solution and vice versa. For instance, consider the combination of aqueous solutions of sodium chloride and silver nitrate. The molecular equation for the reaction is given below:



In the above example, note that when Ag^+ from the silver nitrate solution combines with the Cl^- from the sodium chloride solution, a precipitate of AgCl is formed. Hence, this metathesis reaction is an example of a precipitation reaction.

The fact that AgCl will be a precipitate in this reaction can be predicted by using the "Solubility Rules". The reactions studied in this experiment will help determine some of these solubility rules.

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