

Balancing Chemical Equations Worksheet

- $2 \text{ H}_2 + \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$
- $\text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
- $\text{ S}_8 + 12 \text{ O}_2 \rightarrow 8 \text{ SO}_3$
- $2 \text{ N}_2 + \text{ O}_2 \rightarrow 2 \text{ N}_2\text{O}$
- $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{ O}_2$
- $6 \text{ CO}_2 + 6 \text{ H}_2\text{O} \rightarrow \text{ C}_6\text{H}_{12}\text{O}_6 + 6 \text{ O}_2$
- $\text{ Zn} + 2 \text{ HCl} \rightarrow \text{ ZnCl}_2 + \text{ H}_2$
- $\text{ SiCl}_4 + 4 \text{ H}_2\text{O} \rightarrow \text{ H}_4\text{SiO}_4 + 4 \text{ HCl}$
- $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + \text{ H}_2$
- $2 \text{ H}_3\text{PO}_4 \rightarrow \text{ H}_4\text{P}_2\text{O}_7 + \text{ H}_2\text{O}$
- $\text{ C}_{10}\text{H}_{16} + 8 \text{ Cl}_2 \rightarrow 10 \text{ C} + 16 \text{ HCl}$
- $\text{ CO}_2 + 2 \text{ NH}_3 \rightarrow \text{ OC}(\text{NH}_2)_2 + \text{ H}_2\text{O}$
- $4 \text{ Si}_2\text{H}_3 + 17 \text{ O}_2 \rightarrow 8 \text{ SiO}_2 + 6 \text{ H}_2\text{O}_3$
- $2 \text{ Al}(\text{OH})_3 + 3 \text{ H}_2\text{SO}_4 \rightarrow \text{ Al}_2(\text{SO}_4)_3 + 6 \text{ H}_2\text{O}$
- $4 \text{ Fe} + 3 \text{ O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3$
- $\text{ Fe}_2(\text{SO}_4)_3 + 6 \text{ KOH} \rightarrow 3 \text{ K}_2\text{SO}_4 + 2 \text{ Fe}(\text{OH})_3$
- $2 \text{ C}_7\text{H}_6\text{O}_2 + 15 \text{ O}_2 \rightarrow 14 \text{ CO}_2 + 6 \text{ H}_2\text{O}$
- $\text{ H}_2\text{SO}_4 + 8 \text{ HI} \rightarrow \text{ H}_2\text{S} + 4 \text{ I}_2 + 4 \text{ H}_2\text{O}$
- $4 \text{ FeS}_2 + 11 \text{ O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3 + 8 \text{ SO}_2$
- $2 \text{ Al} + 3 \text{ FeO} \rightarrow \text{ Al}_2\text{O}_3 + 3 \text{ Fe}$
- $\text{ Fe}_2\text{O}_3 + 3 \text{ H}_2 \rightarrow 2 \text{ Fe} + 3 \text{ H}_2\text{O}$
- $\text{ Na}_2\text{CO}_3 + 2 \text{ HCl} \rightarrow 2 \text{ NaCl} + \text{ H}_2\text{O} + \text{ CO}_2$
- $2 \text{ K} + \text{ Br}_2 \rightarrow 2 \text{ KBr}$
- $\text{ C}_7\text{H}_{16} + 11 \text{ O}_2 \rightarrow 7 \text{ CO}_2 + 8 \text{ H}_2\text{O}$
- $\text{ P}_4 + 5 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_5$

Chemical Equations Reactions Section 2 Answers

CH Cherryholmes



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