

Mastery Properties and Changes Test**Multiple Choice**

Choose the choice that best completes the statement or answers the question.

Choose the item that is an example of the term.

- _____ 1. compound
 a. salt
 b. air
 c. metal
 d. mixture
- _____ 2. homogeneous mixture
 a. oil and vinegar
 b. air
 c. metal
 d. mixture
- _____ 3. element
 a. salt
 b. brass
 c. aluminum
 d. mixture

Matching

Match the terms below with their correct definitions.

- | | |
|--------------------------------|---------------------|
| a. law of multiple proportions | g. crystallization |
| b. percent by mass | h. chromatography |
| c. periodic table | i. states of matter |
| d. law of conservation of mass | j. solid |
| e. law of definite proportions | k. liquid |
| f. solution | |
- _____ 4. A homogeneous mixture
- _____ 5. States that mass is neither created nor destroyed in a chemical reaction
- _____ 6. States that regardless of the amount, a compound is always composed of the same elements in the same proportion by mass
- _____ 7. States that different masses of one element combine with the same relative mass of another element to form different compounds
- _____ 8. The grid organizing elements into periods and groups
- _____ 9. The ratio of the mass of each element to the total mass of the compound
- _____ 10. Three physical forms that describe all matter that exists on Earth
- _____ 11. A separation technique that results in the formation of pure solid particles from a solution
- _____ 12. A technique that separates the components of a mixture
- _____ 13. A form of matter with a constant volume that takes the shape of its container

Chemistry Matter And Change Test

Angelica M. Stacy



Chemistry Matter And Change Test:

Glencoe Chemistry Standardized Test Practice McGraw-Hill/Glencoe, 2007-05 **Standardized Test Practice: Student Edition - Glencoe Chemistry Matter and Change**, 2004 **GED Exam Prep 2026-2027** Bill T Reese, 2026-01-04

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SAT Subject Test Chemistry Joseph A. Mascetta, Mark Kernion, 2020-12-01

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Merrill Laboratory Chemistry L. Neal Carmichael, David F. Haines, 1990-10

Laboratory Manual of Chemistry James E. Armstrong, James H. Norton, 1891

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Tests in Chemistry Charles Elwood Dull, 1958

Laboratory Work in Chemistry Edward Harrison Keiser, 1895

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