

Energy and Chemical Change

Section 15.1 Energy

pages 516–522

Practice Problems

pages 519–521

1. A fruit and oatmeal bar contains 142 nutritional Calories. Convert this energy to calories.

$$142 \text{ Calories} = 142 \text{ kcal}$$

$$142 \text{ kcal} \times \frac{1000 \text{ cal}}{1 \text{ kcal}} = 142,000 \text{ cal}$$

2. An exothermic reaction releases 86.5 kJ. How many kilocalories of energy are released?

$$86.5 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = 20.7 \text{ kcal}$$

3. **Challenge** Define a new energy unit, named after yourself, with a magnitude of one-tenth of a calorie. What conversion factors relate this new unit to joules? To Calories?

$$\text{Unit X} = 0.1 \text{ cal}$$

$$1 \text{ cal} = 4.184 \text{ J}$$

$$\text{X} = (0.1 \text{ cal})(4.184 \text{ J/cal}) = 0.4184 \text{ J}$$

$$1 \text{ cal} = 0.001 \text{ Calorie}$$

$$\text{X} = (0.1 \text{ cal})(1 \text{ Cal}/1000 \text{ cal}) = 0.0001 \text{ Calorie}$$

4. If the temperature of 34.4 g of ethanol increases from 25.0°C to 78.8°C, how much heat has been absorbed by the ethanol? Refer to Table 15.2.

$$q = c \times m \times \Delta T$$

$$q = 2.44 \text{ J/(g}\cdot\text{°C)} \times 34.4 \text{ g} \times 53.8\text{°C} = 4.52 \times 10^3 \text{ J}$$

5. A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance? Identify the substance among those listed in Table 15.2 on page 520.

$$q = c \times m \times \Delta T$$

$$c = \frac{q}{m\Delta T} = \frac{(5696 \text{ J})}{(155 \text{ g})(40.0 - 25.0\text{°C})} = 2.45 \text{ J/(g}\cdot\text{°C)}$$

The specific heat is very close to the value for ethanol.

6. **Challenge** A 4.50-g nugget of pure gold absorbed 276 J of heat. The initial temperature was 25.0°C. What was the final temperature?

$$q = c \times m \times \Delta T$$

$$\Delta T = \frac{q}{cm} = \frac{(276 \text{ J})}{(0.129 \text{ J/g}\cdot\text{°C})(4.50 \text{ g})} = 475\text{°C}$$

$$\Delta T = T_f - T_i = T_f - 25.0\text{°C} = 475\text{°C}$$

$$T_f = 5.00 \times 10^2\text{°C}$$

Section 15.1 Assessment

page 522

7. **Explain** how energy changes from one form to another in an exothermic reaction. In an endothermic reaction.

Chemical potential energy changes to heat in exothermic reactions and the heat is released. In endothermic reactions, heat is absorbed and changed to chemical potential energy.

8. **Distinguish** between kinetic and potential energy in the following examples: two separated magnets; an avalanche of snow; books on library shelves; a mountain stream; a stock-car race; separation of charge in a battery.

Two separated magnets illustrate potential energy. In a snow avalanche, positional potential energy is changing to kinetic energy. Books on a shelf illustrate positional potential energy. As water races down a mountain stream, positional potential energy is changing to kinetic energy. In a stock-car race, chemical potential energy is being changed to kinetic energy. The separation of charge in a battery illustrates electrical potential energy.

9. **Explain** how the light and heat of a burning candle are related to chemical potential energy.

Chemical potential energy, contained in the candle, is changed to energy in the form of light and heat and released as the chemical combustion reaction takes place.

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Silberberg



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Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools

Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change

Chemistry: Molecules, Matter, and Change Media Activities Book Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs

Chemistry McGraw-Hill Staff, 2001-03

Chemistry: Science of Matter, Energy, and Change Gregory R. Choppin, Bernard Jaffe, 1965

Chemistry James E. Brady, John R. Holum, 1993-01-21 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

[A Course in inorganic chemistry for colleges](#) Lyman Churchill Newell, 1909

Elementary Chemistry Robert Hart Bradbury, 1903

Journal of the American Chemical Society American Chemical Society, 1928 Proceedings of the Society are included in v 1 59 1879 1937

Prentice Hall Chemistry, 2000

Chemistry E. Russell Hardwick, 1965

Chemical News and Journal of Industrial Science, 1902

Chemistry Henry Dorin, 1987

[The Chemical News and Journal of Physical Science](#), 1911

Journal of the Society of Chemical Industry Society of Chemical Industry (Great Britain), 1888 Includes list of members 1882 1902 and proceedings of the annual meetings and various supplements

Teaching High School Science Alexander Joseph, 1961

Norman Hall's Asvab Preparation Book Norman Hall, 2015-01-02 Everything you need to know thoroughly covered in one book five ASVAB practice tests answer keys tips to boost scores military enlistment information study aids

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Chemical News and Journal of Physical Science, 1898

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Table of Contents Chemistry Matter Change Chapter 15 Answer Key

1. Understanding the eBook Chemistry Matter Change Chapter 15 Answer Key
 - The Rise of Digital Reading Chemistry Matter Change Chapter 15 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter Change Chapter 15 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Matter Change Chapter 15 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter Change Chapter 15 Answer Key
 - Personalized Recommendations
 - Chemistry Matter Change Chapter 15 Answer Key User Reviews and Ratings
 - Chemistry Matter Change Chapter 15 Answer Key and Bestseller Lists
5. Accessing Chemistry Matter Change Chapter 15 Answer Key Free and Paid eBooks
 - Chemistry Matter Change Chapter 15 Answer Key Public Domain eBooks
 - Chemistry Matter Change Chapter 15 Answer Key eBook Subscription Services
 - Chemistry Matter Change Chapter 15 Answer Key Budget-Friendly Options
6. Navigating Chemistry Matter Change Chapter 15 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Matter Change Chapter 15 Answer Key Compatibility with Devices
 - Chemistry Matter Change Chapter 15 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Matter Change Chapter 15 Answer Key
 - Highlighting and Note-Taking Chemistry Matter Change Chapter 15 Answer Key
 - Interactive Elements Chemistry Matter Change Chapter 15 Answer Key

8. Staying Engaged with Chemistry Matter Change Chapter 15 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter Change Chapter 15 Answer Key
9. Balancing eBooks and Physical Books Chemistry Matter Change Chapter 15 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter Change Chapter 15 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Matter Change Chapter 15 Answer Key
 - Setting Reading Goals Chemistry Matter Change Chapter 15 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Matter Change Chapter 15 Answer Key
 - Fact-Checking eBook Content of Chemistry Matter Change Chapter 15 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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