

Electron Configuration Review Test

Electron configuration can be written for each element by following the order provided in the table below in the order of increasing energy.

1. Which element is the primary reactant in all organic combustion and is classified as being unreactive?

- A. Oxygen (O)
- B. Nitrogen (N)
- C. Carbon (C)
- D. Hydrogen (H)

2. Which is the electron configuration of a neutral oxygen atom in its ground state?

- A. $1s^2$
- B. $1s^2 2s^2$
- C. $1s^2 2s^2 2p^4$
- D. $1s^2 2s^2 2p^6$

3. Which of the following elements is a noble gas and is highly unreactive?

- A. Fluorine (F)
- B. Neon (Ne)
- C. Phosphorus (P)
- D. Argon (Ar)

4. Which is the electron configuration for sodium, a primary component of table salt?

- A. $1s^2$
- B. $1s^2 2s^2$
- C. $1s^2 2s^2 2p^6$
- D. $1s^2 2s^2 2p^6 3s^1$

5. Which element's participation in a reaction results in a change from the neutral to an ionic state?

- A. Fluorine
- B. Neon
- C. Phosphorus
- D. Argon

6. Which is the electron configuration for potassium, used as a component of many fertilizer blends?

- A. $1s^2 2s^2 2p^6$
- B. $1s^2 2s^2 2p^6 3s^1$
- C. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- D. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

Chemistry Multiple Choice Test Questions And Answers

Arshad Iqbal



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