

3	A little of the sample is treated with <b>dil. HCl acid</b>	Colourless, odourless gas ( $\text{CO}_2$ ) is evolved with effervescence.	Presence of $\text{CO}_3^{2-}$ ion
4	A little of the sample is treated with <b>con. <math>\text{H}_2\text{SO}_4</math></b> .	White fuming, pungent smelling gas (HCl) produces dense white fumes when mouth of $\text{NH}_4\text{OH}$ bottle is shown near the mouth of the test tube. <b>Note: Warm/Heat if necessary</b>	Presence of $\text{Cl}^-$ ion
5	A pinch of salt is heated with a few drops of conc. $\text{H}_2\text{SO}_4$ for 10-15 seconds.	Reddish brown gas ( $\text{NO}_2$ ) on <b>boiling</b> the mixture with copper turnings and a few drops of <b>conc. <math>\text{H}_2\text{SO}_4</math></b>	Presence of $\text{NO}_3^-$ ion.

### CONFIRMATORY TESTS FOR ANIONS

6	An aqueous solution of the sample is acidified with <b>dil.HCl</b> followed by the addition of <b><math>\text{BaCl}_2</math></b> solution.	A white precipitate ( $\text{BaSO}_4$ ) formed, which is insoluble in <b>con.HCl</b>	Presence of $\text{SO}_4^{2-}$ ion
7	An aqueous solution of the sample is acidified with dilute acetic acid followed by the addition of lead acetate [ $\text{CH}_3\text{-$	A white precipitate ( $\text{PbSO}_4$ ) formed, which is soluble in Ammonium acetate [ $\text{CH}_3\text{COONH}_4$ ] solution.	$\text{SO}_4^{2-}$ ion is confirmed.

# Chemistry Practicals For Class 11 Salt Analysis

**Mark Steyvers**



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