



CHEMISTRY

PERIODIC TABLE OF THE ELEMENTS

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra	Ac	Rf	Mn	Co	Ni	Pd	Ag	Cd	In	Tl	Pb	Bi	Po	At	Rn	

PERIODIC TRENDS

- Atomic Size**
 - Increases from top-left to bottom-right.
 - Increases from top to bottom in a group.
 - Decreases from left to right in a period.
- Ionization Energy**
 - Increases from bottom-left to top-right.
 - Increases from left to right in a period.
 - Decreases from top to bottom in a group.
- Electronegativity**
 - Increases from bottom-left to top-right.
 - Increases from left to right in a period.
 - Decreases from top to bottom in a group.

ATOMIC STRUCTURE AND THE PERIODIC TABLE

- Atomic Structure**
 - Atoms are made of protons, neutrons, and electrons.
 - Electrons are arranged in shells.
 - The number of shells is equal to the period number.
 - The number of electrons in the outer shell is equal to the group number.

$1 - 2 - 8 - 18 - 32$

where, $n = 1, 2, 3, 4, 5$ and so on.

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IONIZATION ENERGY AND ELECTRONEGATIVITY



Ionization energy is the energy required to remove an electron from an atom in its ground state.

Electronegativity is the ability of an atom to attract electrons towards itself.

RELATIVE SIZES OF ATOMS

- The size of an atom increases from top-left to bottom-right.
- The size of an atom increases from top to bottom in a group.
- The size of an atom decreases from left to right in a period.
- The size of an atom is directly proportional to the square root of the atomic number.



THE WAVE MECHANICAL MODEL

- The wave mechanical model describes the behavior of electrons in an atom as waves.
- The energy levels are quantized.
- The probability of finding an electron in a particular region is given by the square of the wave function.

- The energy of an electron in an atom is quantized.
- The energy levels are given by the equation: $E_n = -\frac{13.6}{n^2}$ eV.
- The energy difference between two levels is given by: $\Delta E = E_2 - E_1$.
- The wavelength of light emitted or absorbed is given by: $\lambda = \frac{hc}{\Delta E}$.

PERIODS

- There are seven periods in the periodic table.
- The first period has two elements.
- The second and third periods have eight elements each.
- The fourth and fifth periods have 18 elements each.
- The sixth and seventh periods have 32 elements each.

PERIODIC TABLE

- The periodic table is a table of elements arranged in rows and columns.
- The elements are arranged in order of increasing atomic number.
- The periodic table is divided into groups and periods.
- The groups are numbered 1 to 18.
- The periods are numbered 1 to 7.
- The periodic table shows the trends in atomic size, ionization energy, and electronegativity.
- The periodic table is used to predict the properties of elements.
- The periodic table is a key tool in chemistry.

WORKED EXAMPLES

- Example 1:** Calculate the number of protons, neutrons, and electrons in an atom of an element with atomic number 11 and mass number 23.
- Example 2:** Calculate the number of protons, neutrons, and electrons in an ion of an element with atomic number 16 and mass number 32, which has a charge of 2-.
- Example 3:** Calculate the number of protons, neutrons, and electrons in an atom of an element with atomic number 19 and mass number 39.

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