

Chemistry Unit 1 Review Sheet

The History of the Atom

You don't need to memorize any specifics about the history or the scientists. However, you may be asked to look at some information about a scientist and/or an experiment and identify what major discovery was made. For example, JJ Thomson's cathode ray tube experiment helped identify the Electron using the idea of Charge attraction and repulsion

Vocabulary

Matter

Atom

Nucleus

Proton

Neutron

Electron

Element

Atomic Number

Mass Number

Ion

Net Charge

AMU

Orbital

Wave Mechanical Model

Bohr Model

Electron Configuration

Ground State

Excited State

Bright Line Spectrum

Isotope

Atomic Mass



$$\text{Mass Number} = p^+ + n^0$$

*Electrons DON'T Count!

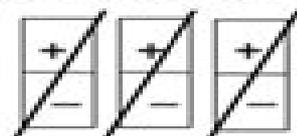
Nuclear Symbols Mass # \rightarrow 12 **C**
 Atomic # \rightarrow 6

Proton	1 amu
Neutron	1 amu
Electron	0 amu

Ions and Net Charge



Neutral Atoms have EQUAL p^+ & e^-



$$(3p^+) \text{ \& \; } (3e^-)$$

$$(+3) + (-3)$$

$$3 - 3 = 0$$



IONS DO NOT have same # p^+ & e^-



$$(3p^+) \text{ \& \; } (2e^-)$$

$$(+3) + (-2)$$

$$3 - 2 = +1$$



**Positive Charge \rightarrow More p^+ **
 **Negative Charge \rightarrow More e^- **

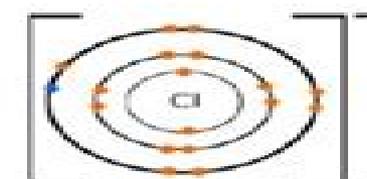
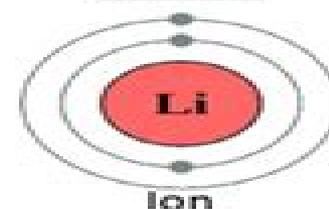
Bohr Models

1st Shell = 2 e^-

2nd Shell = 8 e^-

3rd Shell = *8* e^-

Neutral



Electron Configurations

Neutral Ground State \rightarrow Ca: 2 - 8 - 8 - 2

Ionic Ground State \rightarrow Ca²⁺: 2 - 8 - 8

Excited State \rightarrow Ca: 2 - 8 - 7 - 3

Excited Electrons

- Electrons closer to the nucleus have LOWER energy.
- Electrons further from the nucleus have HIGHER energy.
- An Excited Electrons GAINS energy and Jumps to a higher energy level.
- A De-Excited Electrons RELEASES energy and falls back to ground state.

\rightarrow THIS IS WHAT PRODUCES THE BRIGHT LINE SPECTRUM



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